



Lab Access Card

Classifying Chemical Reactions

Directions: Write the type of chemical reaction in the space provided next to each equation. Name the product or products in the space provided at the end of the equation.

Column I

1. _____ 1. A precipitate forms in a liquid solution containing these reactants. *(magnesium chloride + potassium hydroxide)*
2. _____ 2. Two elements undergo oxidation-reduction reaction.
3. _____ 3. A hydrocarbon explodes, resulting in carbon dioxide.
4. _____ 4. Chlorine atoms break down into two chlorine atoms.
5. _____ 5. A type of combustion reaction that produces heat and light.

Directions: Write the chemical formula for each of the products formed in the space provided.

- combustion _____ 6. $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- combustion _____ 7. $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

combustion

synthesis reaction

- _____ 8. $\text{Mg} + \text{Cl}_2 \rightarrow \text{MgCl}_2$
- _____ 9. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

synthesis reaction

- _____ 10. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
- _____ 11. $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

combustion

single displacement reaction

- _____ 12. $\text{Zn} + \text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$

single displacement reaction

single displacement reaction

- _____ 13. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

combustion

single displacement reaction

- _____ 14. $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

single displacement reaction

- _____ 15. $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

single displacement reaction

Column II

1. synthesis reaction
2. decomposition reaction
3. combustion
4. single displacement reaction
5. double displacement reaction

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Section 2 Reinforcement Chemical Properties

Thomas Griffiths



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