

# Putting Ions in Their Hands

## Covalent, Ionic and Metallic Bonding



### Introduction

The first time seeing formulas and names of ionic compounds can be confusing. However, upon further investigation, there are actually simplistic formulated steps to do so. Discover a systematic method of writing formulas for ionic compounds without the need to experimentally determine each formula using cation and anion cards.

### Concepts

- Ionic compounds
- Formula writing
- Cations and anions

### Background

The *chemical formula* of a compound indicates the kinds of atoms and the relative number of each atom that is combined together to form the compound. *Subscripts* are the numbers written after the chemical element symbols and are used to indicate the number of atoms of each element in the compound. Water, the most abundant chemical on Earth, has the chemical formula  $\text{H}_2\text{O}$  and has a whole-number hydrogen:oxygen atom ratio of 2 to 1.

Some compounds such as sodium chloride, which is common table salt, are composed of *ions*. Ions are atoms or groups of atoms that have a net positive or negative charge. *Cations* are atoms or groups of atoms with a net positive charge, such as  $\text{Ca}^{2+}$  or  $\text{NH}_4^+$ , while *anions* are atoms or groups of atoms with a net negative charge, such as  $\text{Cl}^-$  or  $\text{NO}_3^-$ . Atoms of metallic elements tend to form cations while atoms of nonmetallic elements tend to form anions. Ions join together because of an oppositely-charged or electrostatic attraction between a positive cation and a negative anion resulting in an *ionic bond*. Compounds held together by ionic bonds are *ionic compounds*.

Ionic compounds are generally composed of an extended network of positively-charged metal ions and negatively-charged nonmetal ions. Although composed of ions, ionic compounds are electrically neutral, i.e., the total positive charge is equal to the total negative charge. The chemical formula indicates the smallest whole-number ratio of each element in the network of ions and is called a *formula unit*. For example, magnesium chloride has the chemical formula  $\text{MgCl}_2$ . The magnesium cation ( $\text{Mg}^{2+}$ ) and chloride anions ( $\text{Cl}^-$ ) combine in a 1:2 ratio to form  $\text{MgCl}_2$ . The overall charge on the resulting ionic compound is zero. Thus the final compound is composed of two  $\text{Cl}^-$  ions for each  $\text{Mg}^{2+}$  ion.

Some ions consist of a group of covalently bonded atoms that tend to stay together as if they were a single ion. Such ions are called *polyatomic ions*. An example is the nitrate ion,  $\text{NO}_3^-$ . This polyatomic anion contains one nitrogen atom covalently bonded to three oxygen atoms and has an overall charge of  $-1$ . Calcium cations,  $\text{Ca}^{2+}$ , in the presence of nitrate anions,  $\text{NO}_3^-$ , will combine in a ratio of two nitrate ions for every calcium ion in order to balance the positive and negative charges and achieve electrical neutrality. Thus, the formula for calcium nitrate is written as  $\text{Ca}(\text{NO}_3)_2$ , showing that for every calcium ion there are two nitrate ions. Notice that the polyatomic ion is put in parentheses as a unit with the subscript on the outside of the parentheses.

Writing formulas for ionic compounds may be simply determined by following the steps below:

1. The cation is always written first; the anion is written second.
2. The overall charge for an ionic compound is always zero. The charges of the individual ions are never included in the compound's formula.
3. Subscripts are used to indicate the number of atoms of each element in the compound.
4. Parentheses are used around polyatomic ions to show that the subscript pertains to the polyatomic ion as a whole.
5. Assumed subscripts of one are omitted when writing chemical formulas.

### Materials

Data Table

Envelope

Ion Formula Chart

Scissors

Sheet of anions

Sheet of cations

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**Samuel Johnson**



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