

Module 8: Gas Gases

1. Dalton's law states that the total pressure exerted by a mixture of gases is equal to the sum of the partial pressures exerted by each individual gas. Dalton's law is based on the fact that the total pressure exerted by a mixture of gases is equal to the sum of the partial pressures exerted by each individual gas.

- a. Dalton's law
b. Boyle's law
c. Charles's law
d. Avogadro's law

Dalton's Law (Open Access)		
Gas	Partial Pressure	Total Pressure
Gas 1	1 atm	
Gas 2	2 atm	
Gas 3	3 atm	
Gas 4	4 atm	
Gas 5	5 atm	

2. Dalton's law states that the total pressure exerted by a mixture of gases is equal to the sum of the partial pressures exerted by each individual gas. Dalton's law is based on the fact that the total pressure exerted by a mixture of gases is equal to the sum of the partial pressures exerted by each individual gas.

- a. Why does the total pressure of a mixture of gases increase with temperature?

Because the molecules move faster and exert more pressure on the walls of the container.



- b. Why is the total pressure of a mixture of gases higher than the pressure of each individual gas?

Pogil Activity4 Gas Laws And Temperature Scalesanswers

Lei Huang



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