Name:\_\_\_\_\_\_ Date: \_\_\_\_\_ Period:\_\_\_\_\_

**POGIL- Stoichiometry** 

How do chemists use balanced chemical equations?

What happened to Avogadro when he got bit by 6.02 x 10 23 mosquitoes? He got Mol-aria

#### Why?

Chemists use balanced chemical equations as a basis to calculate how much reactant is needed or product is formed in a reaction. This is called <u>Stoichiometry-(stoi-key-ah-meh-tree)</u> Another way of looking at it is using the <u>mole ratio</u> from the balanced equation and information about one compound in the reaction to determine information about another compound in the equation. A <u>mole ratio</u> is a conversion factor <u>derived from the coefficients of a balanced chemical equation</u> interpreted in terms of moles. In chemical calculations, mole ratios are used to convert between moles one thing and moles of another.

Purpose: In this activity we will address the question: How do I convert between different chemical species in a given reaction?

# MODEL 1: Mole-to-Mole Stoichiometry (1-step)

## Example: mole-to-mole

**Question:** Given the following equation:

 $1 N_2(g) + 3 H_2(g) \rightarrow 2 NH_3(g)$ 

How many moles of H2 are needed to produce 34.8 moles of NH1?

Answer:

Mole ratio

34.8 mol NH<sub>3</sub> | 3 mol H<sub>2</sub> = 52.2mol H<sub>2</sub>

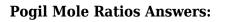
2 mol NH<sub>3</sub>

#### **Key Questions:**

- What was the given in the example above?
- 2. Where did the  $\frac{3 \text{ mot} H_2}{2 \text{ mot} N H_3}$ , come from?
- Double check to see if 52.2 moles of H2 is the correct answer. SHOW the math below.

# **Pogil Mole Ratios Answers**

Michael Rajnik



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## **Table of Contents Pogil Mole Ratios Answers**

- 1. Understanding the eBook Pogil Mole Ratios Answers
  - The Rise of Digital Reading Pogil Mole Ratios Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Pogil Mole Ratios Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pogil Mole Ratios Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Pogil Mole Ratios Answers
  - Personalized Recommendations
  - Pogil Mole Ratios Answers User Reviews and Ratings
  - Pogil Mole Ratios Answers and Bestseller Lists
- 5. Accessing Pogil Mole Ratios Answers Free and Paid eBooks
  - o Pogil Mole Ratios Answers Public Domain eBooks
  - Pogil Mole Ratios Answers eBook Subscription Services
  - Pogil Mole Ratios Answers Budget-Friendly Options

- 6. Navigating Pogil Mole Ratios Answers eBook Formats
  - o ePub, PDF, MOBI, and More
  - Pogil Mole Ratios Answers Compatibility with Devices
  - Pogil Mole Ratios Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Pogil Mole Ratios Answers
  - Highlighting and Note-Taking Pogil Mole Ratios Answers
  - Interactive Elements Pogil Mole Ratios Answers
- 8. Staying Engaged with Pogil Mole Ratios Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Pogil Mole Ratios Answers
- 9. Balancing eBooks and Physical Books Pogil Mole Ratios Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Pogil Mole Ratios Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Pogil Mole Ratios Answers
  - Setting Reading Goals Pogil Mole Ratios Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pogil Mole Ratios Answers
  - Fact-Checking eBook Content of Pogil Mole Ratios Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements

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