

Molecular Geometry

How can molecular shapes be predicted using the VSEPR theory?

Why?

When you draw a Lewis structure for a molecule on paper, you are making a two-dimensional representation of the atoms. In reality, however, molecules are not flat—they are three-dimensional. The 3D shape of a molecule is important because it determines many physical and chemical properties for the substance. In this activity you will learn how to predict molecular shapes.

Model 1 – Lewis Structures

Lewis Structures		3-D Molecular Shape
1. H_2O 	H_2O 2 electron domains (2 bonding, 2 nonbonding)	
2. CO_2 	CO_2 2 electron domains (2 bonding, 0 nonbonding)	
3. CH_4 	CH_4 4 electron domains (4 bonding, 0 nonbonding)	
4. NH_3 	NH_3 4 electron domains (3 bonding, 1 nonbonding)	
5. H_2S 	H_2S 4 electron domains (2 bonding, 2 nonbonding)	
6. Cl_2 	Cl_2 2 electron domains (2 bonding, 0 nonbonding)	

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Pogil Activities For High School Chemistry Molecular Geometry

**Lev Vasil'evich Vilkov, Vladimir
Saidovich Mastroiukov, Nina Ivanovna
Sadova**

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