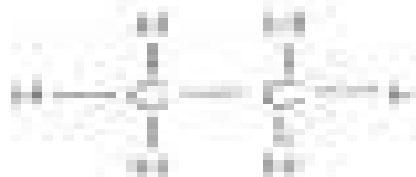


MICHEL 1: Relative electronegativities: $\text{O} > \text{N} > \text{C} > \text{H}$

A



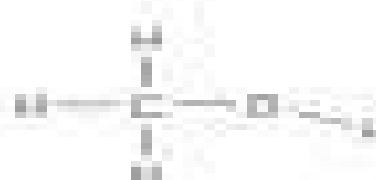
B



C



D



The bonds in B and C are polar.
 The molecules A and D are non-polar.

QUESTIONS

- For each molecule indicated 1, draw in the polar covalent bonds between the atoms. Indicate which electrons are shared and which are not.
 - It may prove to help distinguishing the nature of the bonds that electrons share in the molecule covalent or ionic
 - If both atoms are equally electronegative, draw the electrons in the middle of the bond.
- For each molecule, draw the distribution of all bonding electrons that are shared covalently. If so, write a partial arrangement to show more δ^+ or δ^-
- Label all polar bonds with an arrow (↗).
- Label a polar bond as a dipole moment. Does the δ^+ or δ^- bond polar knowledge or a knowledge for electronegativity. Give the electronegativity values used and give the given K.M.M.1.
- Class molecules that are polar or non-polar or both polar molecules indicated? Is polarisation still given an example?

Pogil Biological Molecules

Lei Huang

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name 'Lei Huang'.

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