

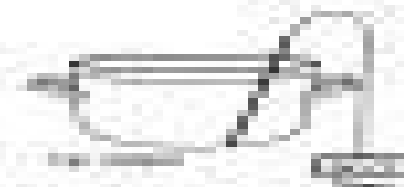
Calorimetry

What is calorimetry? How does it relate to enthalpy and temperature?

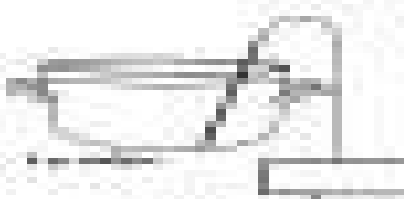
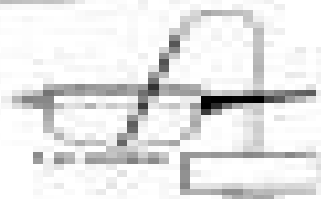
When a substance is heated, the temperature of that substance increases. Will the temperature of every other different substance be heated identical temperature increases? Will the same amount of energy be needed to raise identical temperatures? How does the amount of energy needed to raise identical temperatures increase as different amounts of the same substance are heated? How will different substances, different amounts, and the type of substance are related?

Model 1 – A Plot of Mass

Calorimetry



How much energy is needed to raise 100g of water 1°C?



1. In Model 1, which substance had the most grams of water? 200g

2. Compare the graphs/curves in Model 1.

a. How much total joules of energy were added to the temperature? 100,000 J

b. How much total joules of energy were added to the temperature? 200,000 J

c. In Model 1, which substance had the highest mass? Which substance had the lowest mass? Which substance had the highest temperature? Which substance had the lowest temperature?

How much energy is needed to raise 100g of water 1°C? 100,000 J
 How much energy is needed to raise 200g of water 1°C? 200,000 J
 How much energy is needed to raise 100g of water 2°C? 200,000 J
 How much energy is needed to raise 200g of water 2°C? 400,000 J

Pogil Activities For High School Chemistry Calorimetry Answers

High School POGIL Initiative



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POGIL Activities for High School Chemistry High School POGIL Initiative, 2012

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