

Model 2 – Ternary Acids (Oxyacids)

Acid	Name of Acid in Aqueous Solution	Cation (+)	Polyatomic Anion (-)	Polyatomic Anion Name
HClO_3	Chloric acid	H_3O^+	ClO_3^{-1}	Chlorate
H_2SO_3	Sulfurous acid	$2\text{H}_3\text{O}^+$	SO_3^{-2}	Sulfite
H_2SO_4	Sulfuric acid	$2\text{H}_3\text{O}^+$	SO_4^{-2}	Sulfate
H_3PO_3	Phosphorous acid	$3\text{H}_3\text{O}^+$	PO_3^{-3}	Phosphite
H_3PO_4	Phosphoric acid	$3\text{H}_3\text{O}^+$	PO_4^{-3}	Phosphate
HNO_3	Nitric acid	H_3O^+	NO_3^{-1}	Nitrate
HNO_2	Nitrous acid	H_3O^+	NO_2^{-1}	Nitrite
H_2CO_3	Carbonic acid	$2\text{H}_3\text{O}^+$	CO_3^{-2}	Carbonate

6. Look at the formulas of the ternary acids in Model 2.

a. How are ternary acids different from binary acids in their structure?

Ternary acids contain three elements and they also contain the element oxygen.

b. What number do you think the prefix "ter-" refers to?

"ter" refers to three

7. When ternary acids are mixed with water, ions will form. Fill in the table above with the formulas and names of the anions.

8. Examine the pairs of ternary acids in Model 2 that contain sulfur, phosphorus, and nitrogen. Each pair has one acid that ends in "-ic" and another that ends in "-ous." These endings are related to the name of the polyatomic anion found in the acid ("-ate" or "-ite"). Complete the statements below with the correct acid name ending.

Polyatomic anion ending is "-ate" → acid name ending is "ic"

Polyatomic anion ending is "-ite" → acid name ending is "ous"

9. If the prefix "hydro-" were used to name a ternary acid, what problem would this create when naming HClO_3 ?

ClO_3^{-1} has the name of Chlorate. "ate" must be replaced with "ic". If "hydro" is used as a prefix, it would be hydrochloric acid. This has the formula HCl , not HClO_3 .

10. Write a rule for naming ternary acids.

- The prefix "hydro" is not used in the name when an acid contains more than two elements and one of the elements is oxygen.

- "ate" is replaced with "ic" and "ite" is replaced with "ous".

11. Predict the formula for chlorous acid.



12. Circle the acid(s) below that would be named beginning with the prefix "hydro-."



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1. Understanding the eBook Pogil Activities High School Chemistry Ions Key
 - The Rise of Digital Reading Pogil Activities High School Chemistry Ions Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Activities High School Chemistry Ions Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Activities High School Chemistry Ions Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Activities High School Chemistry Ions Key

- Personalized Recommendations
 - Pogil Activities High School Chemistry Ions Key User Reviews and Ratings
 - Pogil Activities High School Chemistry Ions Key and Bestseller Lists
5. Accessing Pogil Activities High School Chemistry Ions Key Free and Paid eBooks
 - Pogil Activities High School Chemistry Ions Key Public Domain eBooks
 - Pogil Activities High School Chemistry Ions Key eBook Subscription Services
 - Pogil Activities High School Chemistry Ions Key Budget-Friendly Options
 6. Navigating Pogil Activities High School Chemistry Ions Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Pogil Activities High School Chemistry Ions Key Compatibility with Devices
 - Pogil Activities High School Chemistry Ions Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Activities High School Chemistry Ions Key
 - Highlighting and Note-Taking Pogil Activities High School Chemistry Ions Key
 - Interactive Elements Pogil Activities High School Chemistry Ions Key
 8. Staying Engaged with Pogil Activities High School Chemistry Ions Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Activities High School Chemistry Ions Key
 9. Balancing eBooks and Physical Books Pogil Activities High School Chemistry Ions Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Activities High School Chemistry Ions Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Pogil Activities High School Chemistry Ions Key
 - Setting Reading Goals Pogil Activities High School Chemistry Ions Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Pogil Activities High School Chemistry Ions Key

- Fact-Checking eBook Content of Pogil Activities High School Chemistry Ions Key
- Distinguishing Credible Sources

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- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

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- Interactive and Gamified eBooks

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