

## POGIL 2 Ions

How are ions made from neutral atoms?

You have learned that not all atoms of an element are the same. Variation in the number of neutrons results in different isotopes of the element. In this activity we will explore another variation that can take place—the loss and gain of electrons. The exchange of electrons between atoms is a very common way for chemical change to take place. We will see it many times throughout the year.

1. Use Diagram 1 on the next page to complete the following table.

|           | Metal or Nonmetal | Number of protons in the atom | Number of protons in the ion | Number of neutrons in the atom | Number of neutrons in the ion | Number of electrons in the atom | Number of electrons in the ion | Charge on the ion |
|-----------|-------------------|-------------------------------|------------------------------|--------------------------------|-------------------------------|---------------------------------|--------------------------------|-------------------|
| Lithium   | metal             | 3                             | 3                            | 4                              | 4                             | 3                               | 2                              | 1+                |
| Magnesium | <i>Metal</i>      | 12                            | 12                           | 12                             | 12                            | 12                              | 10                             | 2+                |
| Aluminum  | <i>Metalloid</i>  | 13                            | 13                           | 14                             | 14                            | 13                              | 10                             | 3+                |
| Fluorine  | <i>Non metal</i>  | 9                             | 9                            | 10                             | 10                            | 9                               | 10                             | 1-                |
| Oxygen    | nonmetal          | 8                             | 8                            | 8                              | 8                             | 8                               | 10                             | 2-                |
| Nitrogen  | <i>Non metal</i>  | 7                             | 7                            | 7                              | 7                             | 7                               | 10                             | 3-                |

2. Based on the table you completed in Question 1, what distinguishes a neutral atom from an ion?

*A neutral atom has the same number of protons and electrons (no charge). An ion has a charge (# of protons does not equal # of electrons).*

3. Examine the isotope symbols in Diagram 1.

- a. Where is the ion charge located in the isotope symbol?

*Top right corner*

- b. Is a charge indicated on the neutral atoms? If yes, where is it located?

*No*

4. Which subatomic particle carries a positive charge?

*Protons are positive*

5. Which subatomic particle carries a negative charge?

*Electrons are negative*

6. Propose a mathematical equation to calculate the charge on an ion from the number of protons and electrons in an ion. Confirm that your equation works using two positive ion examples and two negative ion examples from Diagram 1.

*The charge of an ion = # of protons - # of electrons*

# Pogil Chemistry Ions Answers

**Marco Cascella**



## POGIL Chemistry Ions Answers:

**Analytical Chemistry** Juliette Lantz, Renée Cole, The POGIL Project, 2014-12-31 An essential guide to inquiry approach instrumental analysis Analytical Chemistry offers an essential guide to inquiry approach instrumental analysis collection The book focuses on more in depth coverage and information about an inquiry approach This authoritative guide reviews the basic principles and techniques Topics covered include method of standard the microscopic view of electrochemistry calculating cell potentials the BerriLambert atomic and molecular absorption processes vibrational modes mass spectra interpretation and much more *Organic Chemistry* Suzanne M. Ruder, The POGIL Project, 2015-12-29 **ORGANIC CHEMISTRY** *Reactions of Ions* William Edward Marsden, 1996 [O Level Chemistry MCQ \(Multiple Choice Questions\)](#) Arshad Iqbal, 2019-06-27 The O Level Chemistry Multiple Choice Questions MCQ Quiz with Answers PDF O Level Chemistry MCQ PDF Download Quiz Questions Chapter 1 14 Practice Tests with Answer Key IGCSE GCSE Chemistry Questions Bank MCQs Notes includes revision guide for problem solving with hundreds of solved MCQs O Level Chemistry MCQ with Answers PDF book covers basic concepts analytical and practical assessment tests O Level Chemistry MCQ PDF book helps to practice test questions from exam prep notes The O Level Chemistry MCQs with Answers PDF eBook includes revision guide with verbal quantitative and analytical past papers solved MCQs O Level Chemistry Multiple Choice Questions and Answers MCQs PDF Free download chapter 1 a book covers solved quiz questions and answers on chapters Acids and bases chemical bonding and structure chemical formulae and equations electricity electricity and chemicals elements compounds mixtures energy from chemicals experimental chemistry methods of purification particles of matter redox reactions salts and identification of ions and gases speed of reaction and structure of atom tests for school and college revision guide O Level Chemistry Quiz Questions and Answers PDF free download eBook s sample covers beginner s solved questions textbook s study notes to practice online tests The book IGCSE GCSE Chemistry MCQs Chapter 1 14 PDF includes high school question papers to review practice tests for exams O Level Chemistry Multiple Choice Questions MCQ with Answers PDF digital edition eBook a study guide with textbook chapters tests for IGCSE NEET MCAT GRE GMAT SAT ACT competitive exam O Level Chemistry Mock Tests Chapter 1 14 eBook covers problem solving exam tests from chemistry textbook and practical eBook chapter wise as Chapter 1 Acids and Bases MCQ Chapter 2 Chemical Bonding and Structure MCQ Chapter 3 Chemical Formulae and Equations MCQ Chapter 4 Electricity MCQ Chapter 5 Electricity and Chemicals MCQ Chapter 6 Elements Compounds and Mixtures MCQ Chapter 7 Energy from Chemicals MCQ Chapter 8 Experimental Chemistry MCQ Chapter 9 Methods of Purification MCQ Chapter 10 Particles of Matter MCQ Chapter 11 Redox Reactions MCQ Chapter 12 Salts and Identification of Ions and Gases MCQ Chapter 13 Speed of Reaction MCQ Chapter 14 Structure of Atom MCQ The Acids and Bases MCQ PDF e Book Chapter 1 practice test to solve MCQ questions on Acid rain acidity needs water acidity or alkalinity acids properties and reactions amphoteric oxides basic acidic neutral and amphoteric chemical formulas chemical reactions

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