

Name: Answer Key Date: \_\_\_\_\_ Period: \_\_\_\_\_  
**POGIL: INTRODUCTION TO CHEMISTRY**

## Model 1: Atomic Structure

*Why?*

Atoms make up everything around us—from the food on our plate to the materials that make up buildings. How do we know what to make those substances out of? How do we know what nutrients are in our food? The answer is that we know about the properties of the atoms that make them up. In order to get to that point, we must first understand the composition of atoms and how to determine that based upon the information provided to us on the periodic table.

| Particle | Symbol | Relative Charge | Absolute Mass              | Relative Mass | Location in the atom |
|----------|--------|-----------------|----------------------------|---------------|----------------------|
| Electron | $e^-$  | -1              | $9.109 \times 10^{-31}$ kg | 0             | Electron Cloud       |
| Proton   | $p^+$  | +1              | $1.673 \times 10^{-27}$ kg | 1             | Nucleus              |
| Neutron  | $n^0$  | 0               | $1.675 \times 10^{-27}$ kg | 1             | Nucleus              |

Memory Clues:

Protons: Positive

3 N's for Neutrons: Neutrons, Neutral, Nucleus

1. A quantum model of the atom is a drawing that represents the nucleus and the electron cloud. Draw the quantum model of the atom, labeling the three subatomic particles described above and labeling the name of their location.



2. What do you notice about the relative mass of the proton and neutron? What about the electron?   
*They are the same mass for the proton and neutron. The mass of the electron is much smaller.*

Since the masses of the subatomic particles are so small, we use another unit to discuss their mass: atomic mass unit (amu). One amu is the mass of a proton. In this way, we can discuss and use the masses of protons (which, relatively speaking, is the same as the mass of a neutron) to determine the atomic mass of a particular atom. The atomic mass is the total number of protons + neutrons in the nucleus of the atom. The atomic mass is the average of all of the masses of the different isotopes of elements (isotopes will be discussed later).

3. Is the mass of an electron taken into account for the atomic mass? Why or why not?   
*NO → it is too small to be significant*
4. 1 amu is equivalent to how many protons?

$$1 \text{ amu} = 1 \text{ proton}$$

# Pogil Atomic Structure

**Edward Uhler Condon**



## **Pogil Atomic Structure:**

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web check your webcam settings and select a higher resolution it seems your web camera streams video as a static image either the webcam hangs or as a video source it uses regular image our test detected a low frame rate fps if you have a high quality webcam make sure that your system is not overloaded and then repeat the test again

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