

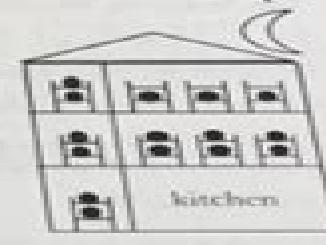
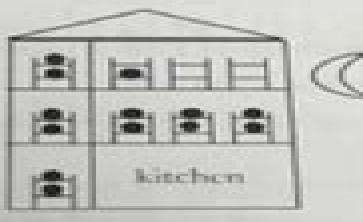
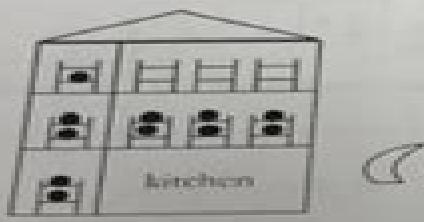
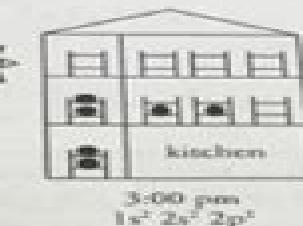
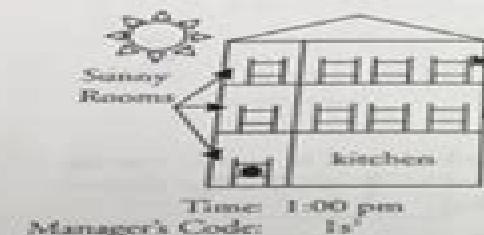
Electron Configurations - POGIL  
Unit 2 Assignment 2

Name: \_\_\_\_\_ Per: \_\_\_\_\_ Date: \_\_\_\_\_

**Learning Target:** I understand that, when an atom is its ground state, that its electrons occupy the lowest-energy levels, sublevels, and orbitals.

**Succesful Criteria:** I can use the periodic table to determine the electron configurations for ground-state atoms.

**Model 1 – The Boarding House**



1. Examine the boarding house diagrams in Model 1. Match each symbol below with its most likely meaning.

a.



I. Bunk bed for boarders

b.



II. Manager's code for the number of boarders in the house and their room assignments.

c.

1s<sup>2</sup>, 2s<sup>2</sup>, 2p<sup>2</sup>, 3s<sup>1</sup>

III. Boarder

# Pogil Activities For High School Chemistry Electron Configurations

**David B. Cook**

## Pogil Activities For High Schol Chemistry Electron Configurations:

Chemistry David John Cole-Hamilton,Iain Patterson,2001    Chemistry Uniy 1 Archie Gibb,Arthur A. Sandison,Learning and Teaching Scotland,Andrew Watson,2000    **Structures and Approximations for Electrons in Molecules** David B.

Cook,1978    **Atomic and Molecular Electronic Configuration Revisited** Harold J. Teague,2014-06-11 In AMEC chapter one you will find a structure for O<sub>2</sub> supported by both bond order and unpaired electron data Previous attempts at providing an electronic structure for O<sub>2</sub> fails due to bond order number of unpaired electrons or oxygen atoms indicating five orbitals although four is the limit Herein you will find the newly described LONE P ORBITAL Pi BOND which is the solution not only for O<sub>2</sub> but also for NO and OF uncharacterized The LPOP bond system is possible due to the small size and high electronegative character of the N O and F atoms which makes it related to H Bonding Diatomic structures involving row two periodic table elements LiBe Be2 BeB B2 BC C2 CN N2 and F2 are also presented All structures are shown being formed by two methods MRAE Most Recently Added Electron and LCAO Linear Combination of Atomic Orbitals Chapter one also includes a Section III MRAE Mechanistic Sequence For NO NO O<sub>2</sub> OF F<sub>2</sub> and F<sub>2</sub> F F<sub>0</sub> Section IV Photographs of NO and O<sub>2</sub> molecular models Section V MRAE MO Energy Level diagrams for the diatomic structures listed above and Section VI a MRAE MO Energy Level Diagrams Template which is based on the cascade of changes in electronic distribution when an electron is forced onto a previous structure MRAE such as N<sub>2</sub> going to N<sub>2</sub> and NO going to NO Chapter 2 includes sections on hybridization and mechanism as it relates to polyatomic species The following sequences are given C H CH H CH<sub>2</sub> H CH<sub>3</sub> H CH<sub>4</sub> saturated and tetrahedral arrangement and N H NH H NH<sub>2</sub> H NH<sub>3</sub> H NH<sub>4</sub> also saturated and tetrahedral arrangement The emphasis is to show trend from atomic structure through sp sp<sub>2</sub> and sp<sub>3</sub> hybridization Chapter 3 is composed of selections from Atomic and Molecular Configuration 1987 Included is MRAE AUFBAU TEMPLATE of the periodic table elements expanded to two pages It is informative to compare the MRAE concept for atoms to the MRAE method for diatomic species in Chapter one In both cases the key is the changes in electronic structure when an electron is added

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bookAnswers are provided for the sections where you need to work out the answers for yourself not the sections where you are just filling things in from a video or website Atoms Electrons Structure and Bonding Workbook Topics Covered are Some of this content has also been published in the Summer Start to A Level Chemistry and a practice exam paperAtomic Structure 20 Multiple choice questionsProperties of Ionic Compounds 15 multiple choice questions Reference table of common ions formulae Formula of Ionic Compounds 65 formulas to work out Drawing Ionic Bonding 10 CompoundsSimple Covalent Bonding 20 multiple choice questionsDrawing Covalent Bonding 10 Compounds Summary Table for the 4 Different Types of Bonding Electron configurationsDrawing electron configurations Drawing Electron Configurations Spot the mistake Electronic Configuration 20 multiple choice questions Exceptions to the Octet Rule Oxidation Numbers 20 multiple choice questions Balancing Equations using the oxidation numbers method 20 to practice Salt Equations 20 equations to complete and balance Shapes of Molecules Investigation Shapes of Molecules and Bond Angles 20 multiple choice questions Electronegativity and Bond Polarity Investigation Intermolecular Bonding 10 multiple choice questions Electrons Structure and Bonding Practice Exam Questions Answers

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