

SOLUBILITY RULES

Soluble:

- All Nitrates, Acetates, Ammonium, and Group 1 (IA) salts
- All Chlorides, Bromides, and Iodides, except Silver, Lead, and Mercury(I)
- All Fluorides except Group 2 (IIA), Lead(II), and Iron(III)
- All Sulfates except Calcium, Strontium, Barium, Mercury, Lead(II), and Silver

Insoluble (0.10 M or greater):

- All Carbonates and Phosphates except Group 1 (IA) and Ammonium
- All Hydroxides except Group 1 (IA), Strontium, Barium, and Ammonium
- All Sulfides except Group 1 (IA), 2 (IIA), and Ammonium
- All Oxides except Group 1 (IA)

Use the solubility rules above to determine if the following compounds would be soluble or insoluble. If the compound is soluble, write (aq) for its state of matter. If the compound is insoluble, write (s).

$\text{Pb}(\text{NO}_3)_2$ () K_2SO_4 () Na_2S ()

CaCl_2 () CaSO_4 () PbS ()

PbCl_2 () KCl () NaNO_3 ()

$\text{Ca}(\text{NO}_3)_2$ () BaCl_2 () Ag_2CO_3 ()

An **insoluble product** that is produced from a double replacement reaction is called a _____

A soluble product that is produced from a double replacement reaction forms our _____

Our net ionic equation will consist of the soluble ions that come together to form our _____

Click on the link: https://javalab.org/en/precipitation_reaction_en/

Reaction 1

Suppose one lead(II) Nitrate reacts with 1 Calcium Chloride. Allow the simulation to run and see what products are produced.

Write the molecular equation that results from this reaction.

___ $\text{Pb}(\text{NO}_3)_2$ (aq) + ___ CaCl_2 (aq) \rightarrow ___ () + ___ ()

Write the net ionic equation for this reaction

___ (aq) + ___ (aq) \rightarrow ___ ()

Pogil Activity Solubility Rules And Net Ionic Equations Answers

DJ Losen



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