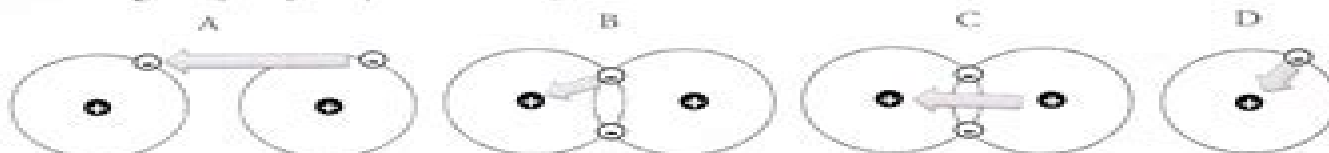


Read This!

Electronegativity is a measure of the ability of an atom's nucleus to attract electrons from a different atom within a covalent bond. A higher electronegativity value correlates to a stronger pull on the electrons in a bond. This value is only theoretical. It cannot be directly measured in the lab.

12. Using the definition stated in the *Read This!* box above, select the best visual representation for electronegativity. Explain your reasoning.



13. Locate the electronegativity values in Model 1.

- a. What is the trend in electronegativity going down a group in Model 1?

Electronegativity decreases as you go down a group (vertical column).

- b. Explain the existence of the trend described in part a in terms of atomic structure and Coulombic attraction.

The ability to attract electrons is based upon attractive force. The greater the distance between the nucleus and valence electrons, the less the attractive force. Less force means a lower electronegativity.

- c. What is the trend in electronegativity going across a period in Model 1?

Electronegativity increases from left to right across a period (horizontal row)

- d. Explain the existence of the trend described in part c in terms of atomic structure and Coulombic attraction.

The greater the number protons in an atom, in a particular period, the greater the attractive force between the nucleus and valence electrons. The greater attractive force means a higher electronegativity.

14. The two diagrams below can summarize each of the three trends discussed in this activity. Write "atomic radius," "ionization energy," and "electronegativity" under the appropriate diagram.



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Ying-Ying Zheng



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Table of Contents Pogil Answer Key Chemistry Periodic Trends

1. Understanding the eBook Pogil Answer Key Chemistry Periodic Trends
 - The Rise of Digital Reading Pogil Answer Key Chemistry Periodic Trends
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Answer Key Chemistry Periodic Trends
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Answer Key Chemistry Periodic Trends
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Answer Key Chemistry Periodic Trends

- Personalized Recommendations
 - Pogil Answer Key Chemistry Periodic Trends User Reviews and Ratings
 - Pogil Answer Key Chemistry Periodic Trends and Bestseller Lists
5. Accessing Pogil Answer Key Chemistry Periodic Trends Free and Paid eBooks
 - Pogil Answer Key Chemistry Periodic Trends Public Domain eBooks
 - Pogil Answer Key Chemistry Periodic Trends eBook Subscription Services
 - Pogil Answer Key Chemistry Periodic Trends Budget-Friendly Options
 6. Navigating Pogil Answer Key Chemistry Periodic Trends eBook Formats
 - ePub, PDF, MOBI, and More
 - Pogil Answer Key Chemistry Periodic Trends Compatibility with Devices
 - Pogil Answer Key Chemistry Periodic Trends Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Answer Key Chemistry Periodic Trends
 - Highlighting and Note-Taking Pogil Answer Key Chemistry Periodic Trends
 - Interactive Elements Pogil Answer Key Chemistry Periodic Trends
 8. Staying Engaged with Pogil Answer Key Chemistry Periodic Trends
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Answer Key Chemistry Periodic Trends
 9. Balancing eBooks and Physical Books Pogil Answer Key Chemistry Periodic Trends
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Answer Key Chemistry Periodic Trends
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Pogil Answer Key Chemistry Periodic Trends
 - Setting Reading Goals Pogil Answer Key Chemistry Periodic Trends
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Pogil Answer Key Chemistry Periodic Trends

- Fact-Checking eBook Content of Pogil Answer Key Chemistry Periodic Trends
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
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14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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