







Part A: Solubility Rules for Ionic Compounds

Use the solubility rules in your textbook or the corresponding lecture to complete question 1.

1) For each of the following ionic compounds in the table below:

- Put an "aq" (for aqueous) in the parentheses after the formula if the compound is soluble in water or put an "s" (for solid) in the parentheses after the formula if the compound is insoluble in water.
- Put in the name of the compound in the space provided.
- Complete the drawing of the beaker of water by showing the compound either broken up into ions (if it is soluble) or sitting on the bottom (if it is insoluble).

The first two drawings (you will need to fill in the sides and names) have been done for you and should be used as examples. Notice that the polyatomic ionic compounds break up into ions that reflect the correct charge and ratios.

Na_2CO_3 () name: _____ 	Cu_2S () name: _____ 	$\text{Pb}(\text{NO}_3)_2$ () name: _____ 
$\text{Mg}(\text{NO}_3)_2$ () name: _____ 	$\text{Al}_2(\text{SO}_4)_3$ () name: _____ 	PbI_2 () name: _____ 

Pogil Answer Key Solubility Rules And Net Ionic Equations

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