

## Model 2 – Ternary Acids (Oxyacids)

| Acid                    | Name of Acid in Aqueous Solution | Cation (+)              | Polyatomic Anion (-) | Polyatomic Anion Name |
|-------------------------|----------------------------------|-------------------------|----------------------|-----------------------|
| $\text{HClO}_3$         | Chloric acid                     | $\text{H}_3\text{O}^+$  | $\text{ClO}_3^{-1}$  | Chlorate              |
| $\text{H}_2\text{SO}_3$ | Sulfurous acid                   | $2\text{H}_3\text{O}^+$ | $\text{SO}_3^{-2}$   | Sulfite               |
| $\text{H}_2\text{SO}_4$ | Sulfuric acid                    | $2\text{H}_3\text{O}^+$ | $\text{SO}_4^{-2}$   | Sulfate               |
| $\text{H}_3\text{PO}_3$ | Phosphorous acid                 | $3\text{H}_3\text{O}^+$ | $\text{PO}_3^{-3}$   | Phosphite             |
| $\text{H}_3\text{PO}_4$ | Phosphoric acid                  | $3\text{H}_3\text{O}^+$ | $\text{PO}_4^{-3}$   | Phosphate             |
| $\text{HNO}_3$          | Nitric acid                      | $\text{H}_3\text{O}^+$  | $\text{NO}_3^{-1}$   | Nitrate               |
| $\text{HNO}_2$          | Nitrous acid                     | $\text{H}_3\text{O}^+$  | $\text{NO}_2^{-1}$   | Nitrite               |
| $\text{H}_2\text{CO}_3$ | Carbonic acid                    | $2\text{H}_3\text{O}^+$ | $\text{CO}_3^{-2}$   | Carbonate             |

6. Look at the formulas of the ternary acids in Model 2.

a. How are ternary acids different from binary acids in their structure?

*Ternary acids contain three elements and they also contain the element oxygen.*

b. What number do you think the prefix "ter-" refers to?

*"ter" refers to three*

7. When ternary acids are mixed with water, ions will form. Fill in the table above with the formulas and names of the anions.

8. Examine the pairs of ternary acids in Model 2 that contain sulfur, phosphorus, and nitrogen. Each pair has one acid that ends in "-ic" and another that ends in "-ous." These endings are related to the name of the polyatomic anion found in the acid ("-ate" or "-ite"). Complete the statements below with the correct acid name ending.

Polyatomic anion ending is "-ate" → acid name ending is "ic"

Polyatomic anion ending is "-ite" → acid name ending is "ous"

9. If the prefix "hydro-" were used to name a ternary acid, what problem would this create when naming  $\text{HClO}_3$ ?

*$\text{ClO}_3^{-1}$  has the name of Chlorate. "ate" must be replaced with "ic". If "hydro" is used as a prefix, it would be hydrochloric acid. This has the formula  $\text{HCl}$ , not  $\text{HClO}_3$ .*

10. Write a rule for naming ternary acids.

*- The prefix "hydro" is not used in the name when an acid contains more than two elements and one of the elements is oxygen.*

*- "ate" is replaced with "ic" and "ite" is replaced with "ous".*

11. Predict the formula for chlorous acid.



12. Circle the acid(s) below that would be named beginning with the prefix "hydro-."



# Pogil Chemistry Polyatomic Answer Key

**Arturo Cuomo**



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## **Table of Contents Pogil Chemistry Polyatomic Answer Key**

1. Understanding the eBook Pogil Chemistry Polyatomic Answer Key
  - The Rise of Digital Reading Pogil Chemistry Polyatomic Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Chemistry Polyatomic Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pogil Chemistry Polyatomic Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Chemistry Polyatomic Answer Key
  - Personalized Recommendations
  - Pogil Chemistry Polyatomic Answer Key User Reviews and Ratings
  - Pogil Chemistry Polyatomic Answer Key and Bestseller Lists
5. Accessing Pogil Chemistry Polyatomic Answer Key Free and Paid eBooks
  - Pogil Chemistry Polyatomic Answer Key Public Domain eBooks
  - Pogil Chemistry Polyatomic Answer Key eBook Subscription Services
  - Pogil Chemistry Polyatomic Answer Key Budget-Friendly Options
6. Navigating Pogil Chemistry Polyatomic Answer Key eBook Formats

- ePub, PDF, MOBI, and More
- Pogil Chemistry Polyatomic Answer Key Compatibility with Devices
- Pogil Chemistry Polyatomic Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Pogil Chemistry Polyatomic Answer Key
  - Highlighting and Note-Taking Pogil Chemistry Polyatomic Answer Key
  - Interactive Elements Pogil Chemistry Polyatomic Answer Key
- 8. Staying Engaged with Pogil Chemistry Polyatomic Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Pogil Chemistry Polyatomic Answer Key
- 9. Balancing eBooks and Physical Books Pogil Chemistry Polyatomic Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Pogil Chemistry Polyatomic Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Pogil Chemistry Polyatomic Answer Key
  - Setting Reading Goals Pogil Chemistry Polyatomic Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pogil Chemistry Polyatomic Answer Key
  - Fact-Checking eBook Content of Pogil Chemistry Polyatomic Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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