

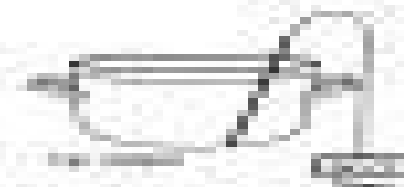
Calorimetry

What is calorimetry? How does it measure heat energy and temperature?

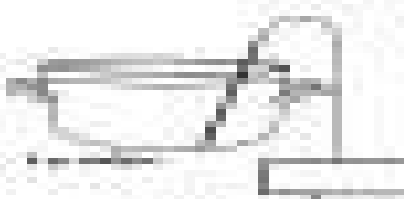
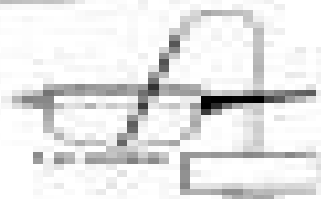
When a substance is heated, the temperature of that substance increases. Will the temperature of every other different substance be heated identical temperature increase? Will the same amount of energy be needed to raise identical temperature increase in different amounts of the same substance? In this practice you will explore how mass, temperature, heat energy, and the type of substance are related.

Model 1 – A Plot of Mass

Calorimetry



How much heat energy is added to the water per minute?



1. In Model 1, which substance had the most grams of water? 100g

2. Compare the graphs/curves in Model 1.

a. How many total joules of energy were added to the substance? 100,000 J

b. How many total joules of energy were added to the substance? 100,000 J

c. In Model 1, which substance had the highest mass? Which substance had the lowest mass? Explain your reasoning.

Why? Explain how mass, temperature, amount of heat energy added, and the type of substance are related. Write formulas for heat added, temperature change, and mass.

Pogil Activities For High School Chemistry Calorimetry Packet

David Baud



Pogil Activities For High School Chemistry Calorimetry Packet:

POGIL Activities for High School Chemistry High School POGIL Initiative, 2012

CHEM-PACS , 1989

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