

Name: _____ Date: _____ Period: _____

What happened to
Avogadro when he
got bit by 6.02×10^{23} mosquitoes?
He got Mol-aria

POGIL- Stoichiometry

How do chemists use balanced chemical equations?

Why?

Chemists use balanced chemical equations as a basis to calculate how much reactant is needed or product is formed in a reaction. This is called **Stoichiometry**- (stoi-key-ah-meh-tree) Another way of looking at it is using the mole ratio from the balanced equation and information about one compound in the reaction to determine information about another compound in the equation. A mole ratio is a conversion factor derived from the coefficients of a balanced chemical equation interpreted in terms of moles. In chemical calculations, mole ratios are used to convert between moles one thing and moles of another.

Purpose: In this activity we will address the question: *How do I convert between different chemical species in a given reaction?*

MODEL 1: Mole-to-Mole Stoichiometry (1-step)

Example: mole-to-mole

Question: Given the following equation:



How many moles of H_2 are needed to produce 34.8 moles of NH_3 ?

Mole ratio

Answer:

$$\frac{34.8 \text{ mol NH}_3}{1} \times \frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3} = 52.2 \text{ mol H}_2$$

Key Questions:

1. What was the given in the example above?
2. Where did the $\frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3}$ come from?
3. Double check to see if 52.2 moles of H_2 is the correct answer. SHOW the math below.

Pogil Chemistry Mole Ratios Answers

Hussin A.Rothana



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Mole Ratios Lifeliqe,2019 This lesson plan relates balanced chemical equations to everyday analogies defines stoichiometry and uses mole ratios to convert between amounts of substances in chemical reactions You Can Do Chemistry Sunil Tanna,2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018 **Counting Moles** Nigel P. Freestone,2016-01-04 Students studying chemistry often struggle with the mole Counting Moles provides an effective aid to learning by giving clear and confident presentation of the essentials of the mole concept needed by those starting chemistry courses This user friendly self teach e book is split into six chapters which sequentially introduce the mole calculating frame to help solve problems Over 200 fully worked examples are given along with several hundred questions The mole concept is applied to topics such as relative atomic mass and relative formula mass percentage composition empirical and molecular formula The book also covers concentration its units volumetric analysis and the relationship between volume mass and moles of gases Counting Moles culminates in you taking a Mole Driving Test On passing this test you are issued with a Counting Moles Driving License that will give you all the confidence required to correctly answer all mole calculations

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