

POGIL: Le Chatelier's Principle

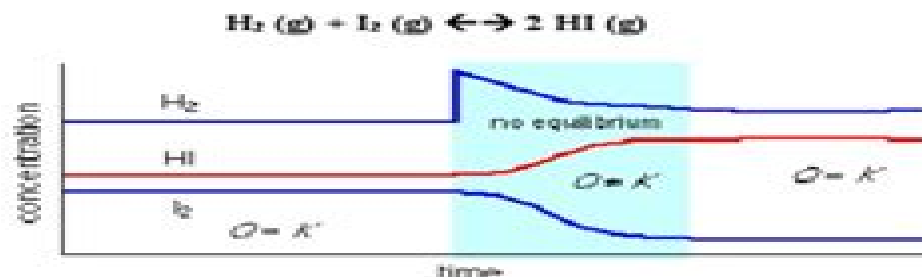
Why?

When opposing forces or issues are balanced, a system is said to be in equilibrium. In biology class you called this "maintaining homeostasis" for living systems or organisms. Placing a stress on any equilibrium system, whether it is chemical, biological, societal, environmental, or personal, causes the equilibrium to be disrupted.

Le Chatelier's Principle allows us to predict the results that follow from "stressing," or in other words, changing the conditions of a system at chemical equilibrium. This allows scientists to develop techniques to control chemical reactions in natural and industrial settings in order to obtain desired products.

MODEL #1: Disrupting an Equilibrium

The graph below represents a sealed system that starts at equilibrium, and is then disrupted when hydrogen gas is injected into the it (increasing H_2 concentration).



According to LeChatelier's principle, the addition of hydrogen gas disrupts the equilibrium system. The graph shows how the system responds to the disturbance of the added hydrogen gas.

QUESTIONS

1. What information on the graph indicates that the system was initially at equilibrium?
2. What information on the graph shows the system was disrupted?
3. By studying the graph, what caused the disruption? (HINT: Look at the concentration information for each substance.)

Pogil Equilibrium And Le Chatelier Answers

Philip J. Hills



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Chemical Equilibria Philip J. Hills,1969 *Le Châtelier's Principle* Lifeliqe,2019 This lesson plan covers how to analyze stresses to a chemical system at equilibrium and predict how it will respond to that stress *Le Chatelier's Principle-Concentration* ,2016 Before you can solve equilibrium problems you need to understand what happens to an equilibrium system when conditions are changed You learn about a fundamental idea Le Chatelier s Principle which lays the groundwork for a broader understanding of equilibrium Chemical Equilibrium Allen J. Bard,1966 **Equilibrium Calculations** John T. Donoghue,1971 **The Strong and Weak Form of the Principle of Le Chatelier-Braun, the Chemical Equilibrium, the Social Equality and Symmetry Law of Pierre Curie** Pekka T. Laakso,1998 *Equilibrium* ,1991 *Equilibrium* Thomas R. Blackburn,1969 *Chemical Equilibrium* Carl John Nyman,Randall E. Hamm,1968 Some Examples of Chemical Equilibria/LeChatelier Principle James M. Postma,Julian Roberts,J. Leland Hollenberg,2000-01-15 *Chemical Equilibrium* Robert Blanchard Fischer,Dennis G. Peters,1970 **On a Theorem of Le Chatelier ...** Paul Saurel,1901* **Principles of Chemical Equilibrium** Kelso B. Morris,1969 **The General Conditions of Validity of the Principle of Le Chatelier** Alfred James Lotka,1922 **Chemical equilibrium ,** *Chemical Equilibrium* Bassam Z. Shakhashiri,1972 HSC Chemical Equilibrium C. M. Roebuck,1989

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