

Name : _____ Date : _____

Percent Composition, Molecular Formula, and Empirical Formula

1. What is the percent composition by mass of all elements in iron (III) oxide?
2. Analysis of a compound's sample indicates that it has 1.04 g K, 0.70 g Cr, and 0.86 g O. What is its empirical formula?
3. If 4.04 g of nitrogen combine with 11.46 g of oxygen to produce a compound with a molar mass of 108 g, what is the molecular formula of this compound? What is its molecular mass?
4. Phosphoric acid is found in some soft drinks. A sample of phosphoric acid contains 3.161 g of phosphorous, 0.3086 g of hydrogen, and 6.531 g of oxygen. What is its empirical formula?
5. What is the empirical formula of a compound with three times as many hydrogen atoms as carbon atoms but only half as many oxygen atoms as carbon atoms?
6. Can the molecular formula of a compound ever be the same as the empirical formula?

Pogil Percent Composition And Empirical Formulas Answers

J Elliott



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