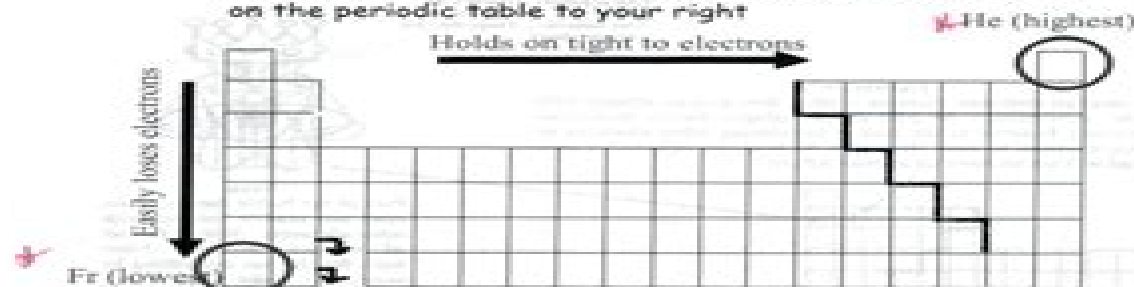


b. Why do elements get larger as you go down a group? Additional energy levels are added to the electron cloud as you go down a group.

Ionization Energy:

1. In simple terms, explain ionization energy. The energy required to remove an electron from a neutral atom.
2. Using the information in figure 5-15 (Pg 143) in your book, write the ionization energy for alkali metals & period #2 on the periodic table to your right



3. On the arrows on the diagram, indicate the trend for ionization energy as you move down a column and across a row. (Options: easily loses an electron OR holds on tight to electrons)
4. Of all the elements on the periodic table, identify & circle the element with the highest & lowest ionization energy.

5. Rank the following elements from highest to lowest ionization energy:

a) Be, Ba, Mg = Be, Mg, Ba

b) Li, F, B = F, B, Li

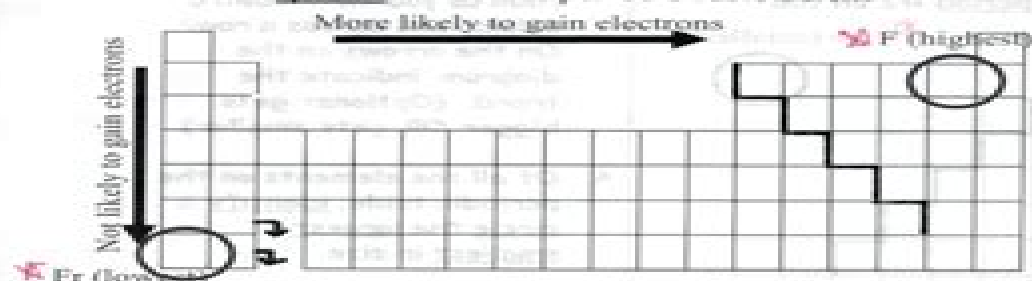
6. Why does it take less energy to remove an electron from a large atom compared to a small atom? Their electrons are closer to the nucleus, so the protons have a larger "pull" on them.

It takes less energy to remove an electron from a large atom because

Electronegativity: the effective nuclear charge decreases and electrons are not held as tightly since

1. Explain Electronegativity: The ability of an atom to get electrons from other atoms.

2. Using the information in figure 5-20 (pg 151) in your book, write the electronegativity for the Halogens and period #2 on the periodic table below.



3. On the arrows on the diagram, indicate the trend for electronegativity as you move down a column and across a row. (Options: more likely to gain electrons, not likely to gain electrons)

4. Of all the elements on the periodic table, identify & circle the element with the highest & lowest electronegativity.

5. Rank the following elements from most to least electronegative element:

a) Be, Ba, Mg = Be, Mg, Ba

b) Li, F, B = F, B, Li

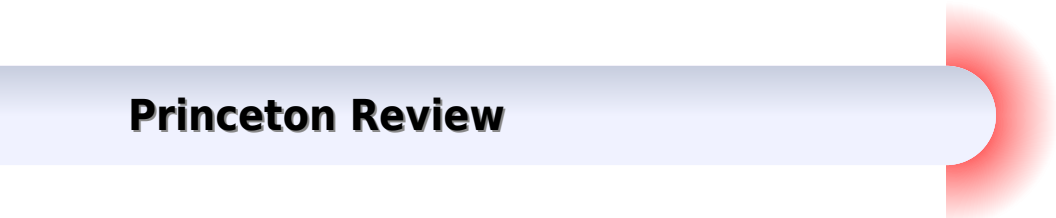
6. Why are small atoms better at pulling electrons off of larger atoms than vice versa?

Their electrons are closer to the nucleus, so the protons have a larger "pull" on the electrons in other atoms.

not detailed enough
The increased number of energy level(s) increases the distance over which nucleus must pull and therefore reduces the attraction for electrons } *AP answer*

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 - The Rise of Digital Reading Pogil Answer Key Chemistry Cracking The
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Answer Key Chemistry Cracking The
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Answer Key Chemistry Cracking The
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Answer Key Chemistry Cracking The
 - Personalized Recommendations
 - Pogil Answer Key Chemistry Cracking The User Reviews and Ratings
 - Pogil Answer Key Chemistry Cracking The and Bestseller Lists
5. Accessing Pogil Answer Key Chemistry Cracking The Free and Paid eBooks
 - Pogil Answer Key Chemistry Cracking The Public Domain eBooks

- Pogil Answer Key Chemistry Cracking The eBook Subscription Services
- Pogil Answer Key Chemistry Cracking The Budget-Friendly Options
- 6. Navigating Pogil Answer Key Chemistry Cracking The eBook Formats
 - ePub, PDF, MOBI, and More
 - Pogil Answer Key Chemistry Cracking The Compatibility with Devices
 - Pogil Answer Key Chemistry Cracking The Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Answer Key Chemistry Cracking The
 - Highlighting and Note-Taking Pogil Answer Key Chemistry Cracking The
 - Interactive Elements Pogil Answer Key Chemistry Cracking The
- 8. Staying Engaged with Pogil Answer Key Chemistry Cracking The
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Answer Key Chemistry Cracking The
- 9. Balancing eBooks and Physical Books Pogil Answer Key Chemistry Cracking The
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Answer Key Chemistry Cracking The
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pogil Answer Key Chemistry Cracking The
 - Setting Reading Goals Pogil Answer Key Chemistry Cracking The
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pogil Answer Key Chemistry Cracking The
 - Fact-Checking eBook Content of Pogil Answer Key Chemistry Cracking The
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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