

## Trends on the Periodic Table

### Part 1: Identifying chemical properties/characteristics of the elements

For each of the following elements, draw a caricature or cartoon above the description to show its "chemical personality" based on the description provided. (Note: See the Neon example for ideas.)

**Lithium**  
Being one of the smaller elements with only 3 electrons, lithium is happy to give one of its electrons away and never seems to ask for any in return.



**Francium**  
Francium is the largest of all of the elements and has so many electrons that it has trouble keeping track of them. Francium will eagerly give up electrons to other elements and never takes in extras. But watch out, it has an extremely reactive and explosive personality!

**Fluorine**  
One of the smallest elements, Fluorine is often described as snappy with a "magnetic" personality. As part of the halogen family (stoutest electron thieves), Fluorine is the best at convincing other elements to give up their electrons to him. However, he holds on tight to his own electrons and will not let them go without a fight.



I am so perfect... I have a full valence shell. I don't need those other elements... I hope they just leave me alone!



**Neon**  
Part of the Noble family, Neon does not readily interact with other elements. With a full set of valence electrons, Neon feels that he has everything he needs to make him happy without sharing with others.

**Iodine**  
As part of the halogen family, Iodine is also a skilled electron thief, however, not nearly as good as its smaller brothers and sisters. Iodine will also put up a pretty good fight to keep its own electrons.

### Part 2: Use your textbook to complete the following:

#### Atomic Radii Size:

1. Explain **atomic radius**: A measure of the size (volume) of an atom. Determined by measuring half of the distance between two identical atoms bonded together.
2. Using the information in figure 5-13 (pg 141) in your book, **write** the atomic radii size for **alkali metals & period #2** on the periodic table to your right.



3. What happens to atomic radii as you move down a column and across a row? On the arrows on the diagram, indicate the trend. (Options: gets bigger OR gets smaller)
4. Of all the elements on the periodic table, **identify & circle the largest & smallest** in size.

5. Rank the following elements from largest to smallest in atomic radii:

a) Be, Ba, Mg = **Ba, Mg, Be**

b) Li, F, B = **Li, B, F**

6. How can you explain the trend:

a. Why do they get smaller as you go across a period? Additional protons in the nucleus "pull" on electrons, bring the electron cloud closer.

*Effective nuclear charge,  $Z_{eff}$ , increases the attraction of the nucleus and therefore pulls the electron cloud closer to the nucleus resulting in smaller radii.*

*not a good enough answer*  
**4AP Answer**

# Pogil Answer Key Periodic Trends Chapter 6

**JA Banks**



## **Pogil Answer Key Periodic Trends Chapter 6:**

## Whispering the Strategies of Language: An Psychological Quest through **Pogil Answer Key Periodic Trends Chapter 6**

In a digitally-driven world where displays reign supreme and instant transmission drowns out the subtleties of language, the profound secrets and psychological nuances concealed within words frequently get unheard. Yet, situated within the pages of **Pogil Answer Key Periodic Trends Chapter 6** a interesting fictional value pulsating with raw feelings, lies a fantastic journey waiting to be undertaken. Published by a talented wordsmith, that marvelous opus attracts visitors on an introspective journey, gently unraveling the veiled truths and profound influence resonating within the very cloth of each word. Within the emotional depths of the poignant review, we will embark upon a heartfelt exploration of the book is core themes, dissect their fascinating publishing design, and fail to the strong resonance it evokes deep within the recesses of readers hearts.

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### **Table of Contents Pogil Answer Key Periodic Trends Chapter 6**

1. Understanding the eBook Pogil Answer Key Periodic Trends Chapter 6
  - The Rise of Digital Reading Pogil Answer Key Periodic Trends Chapter 6
  - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Answer Key Periodic Trends Chapter 6
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pogil Answer Key Periodic Trends Chapter 6
  - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Answer Key Periodic Trends Chapter 6
  - Personalized Recommendations

- Pogil Answer Key Periodic Trends Chapter 6 User Reviews and Ratings
- Pogil Answer Key Periodic Trends Chapter 6 and Bestseller Lists
- 5. Accessing Pogil Answer Key Periodic Trends Chapter 6 Free and Paid eBooks
  - Pogil Answer Key Periodic Trends Chapter 6 Public Domain eBooks
  - Pogil Answer Key Periodic Trends Chapter 6 eBook Subscription Services
  - Pogil Answer Key Periodic Trends Chapter 6 Budget-Friendly Options
- 6. Navigating Pogil Answer Key Periodic Trends Chapter 6 eBook Formats
  - ePub, PDF, MOBI, and More
  - Pogil Answer Key Periodic Trends Chapter 6 Compatibility with Devices
  - Pogil Answer Key Periodic Trends Chapter 6 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Pogil Answer Key Periodic Trends Chapter 6
  - Highlighting and Note-Taking Pogil Answer Key Periodic Trends Chapter 6
  - Interactive Elements Pogil Answer Key Periodic Trends Chapter 6
- 8. Staying Engaged with Pogil Answer Key Periodic Trends Chapter 6
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Pogil Answer Key Periodic Trends Chapter 6
- 9. Balancing eBooks and Physical Books Pogil Answer Key Periodic Trends Chapter 6
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Pogil Answer Key Periodic Trends Chapter 6
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Pogil Answer Key Periodic Trends Chapter 6
  - Setting Reading Goals Pogil Answer Key Periodic Trends Chapter 6
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pogil Answer Key Periodic Trends Chapter 6
  - Fact-Checking eBook Content of Pogil Answer Key Periodic Trends Chapter 6

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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