

14. Provide the chemical formula for each of the following ionic compounds.

Barium chloride $\text{Ba}^{+2} \text{Cl}^{-1} \rightarrow \text{BaCl}_2$ Magnesium oxide $\text{Mg}^{+2} \text{O}^{-2} \rightarrow \text{MgO}$

15. Consider the two chemical formulas you wrote in Question 3 for compounds of iron and sulfur. Would the name "iron sulfide" be sufficient to uniquely identify either of those compounds? Explain.

"iron sulfide" is not sufficient because there are two different forms of iron (Fe^{+2} and Fe^{+3}). The name does not indicate which iron is in the compound.

Read This!

When the metal in an ionic compound always forms an ion with the same charge, you need not indicate that charge as part of the compound name. However, some atoms have the ability to form more than one type of ion. This can make naming confusing. You can't simply refer to a compound of copper and oxygen as "copper oxide." People won't know which compound you are referring to— CuO or Cu_2O .

Model 3 – Ionic Compound Names (Metals that form multiple ions)

Cu_2O Copper(I) oxide	PbO Lead(II) oxide
CuO Copper(II) oxide	PbO_2 Lead(IV) oxide
SnF_2 Tin(II) fluoride	FeCl_2 Iron(II) chloride
SnF_4 Tin(IV) fluoride	FeCl_3 Iron(III) chloride

16. Model 3 is labeled "Metals that form multiple ions." What other metals that form multiple ions could be included in Model 3? Model 1 may be helpful in this regard.

Mercury (Hg_2^{+2} and Hg^{+2}) and nickel (Ni^{+2} and Ni^{+3})

17. Describe the most obvious difference between the names in Model 3 and those in Model 2.

Roman numerals are used in Model 3 and not in Model 2.

18. Do the Roman numerals in the names in Model 3 relate to the number of cations or number of anions in the formula unit? Support your answer by citing two specific examples.

The Roman numerals have no connection between the number of cations or anions in the compound. An example is Lead (II) oxide. There is only one lead (Pb) ion in the compound and not four.

19. Keeping in mind that the sum of the charges in an ionic compound must equal zero, use the chemical formulas in Model 3 to answer the following questions:

- a. Identify the charge on the copper cations in copper(I) oxide and copper(II) oxide, respectively.

$\text{Cu}_2^+ \text{O}^{-2}$ (Each copper I must be 1^+) || $\text{Cu}^{+2} \text{O}^{-2}$ (Each copper II must be 2^+)

- b. Identify the charge on the iron cations in iron(II) chloride and iron(III) chloride, respectively.

$\text{Fe}^{+2} \text{Cl}_2^{-1}$ (Fe^{+2}) || $\text{Fe}^{+3} \text{Cl}_3^{-1}$ (Fe^{+3})

20. What do the Roman numerals in the compounds described in Question 19 indicate?

The Roman numerals indicate the charge of the ion.
(II $\rightarrow 2^+$, III $\rightarrow 3^+$)

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Thomas Griffiths



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