

Model 2 – Proposed Calculations for Mass of NH_3 to Mass of N_2

Toby's Method

$$\frac{x \text{ grams}}{50.0 \text{ g}} = \frac{1 \text{ mole } \text{N}_2}{2 \text{ moles } \text{NH}_3} \rightarrow x = \underline{15} \text{ g } \text{N}_2$$

Rachel's Method

$$50.0 \text{ g } \text{NH}_3 \times \frac{1 \text{ mole } \text{NH}_3}{17.0 \text{ g } \text{NH}_3} = \underline{2.94} \text{ moles } \text{NH}_3$$

$$\frac{x \text{ mole } \text{N}_2}{\underline{1.5} \text{ mole } \text{NH}_3} = \frac{1 \text{ mole } \text{N}_2}{2 \text{ moles } \text{NH}_3} \rightarrow x = \underline{0.75} \text{ moles } \text{N}_2$$

$$\underline{0.75} \text{ mole } \text{N}_2 \times \frac{28.0 \text{ g } \text{N}_2}{1 \text{ mole } \text{N}_2} = \underline{21.0} \text{ g } \text{N}_2$$

Jerry's Method

$$50.0 \text{ g } \text{NH}_3 \times \frac{1 \text{ mole } \text{NH}_3}{17.0 \text{ g } \text{NH}_3} \times \frac{1 \text{ mole } \text{N}_2}{2 \text{ moles } \text{NH}_3} \times \frac{28.0 \text{ g } \text{N}_2}{1 \text{ mole } \text{N}_2} = \underline{21.0} \text{ g } \text{N}_2$$

- Model 2 shows three proposed calculations to solve the problem in Question 10. Complete the calculations in Model 2 by filling in the underlined values.
- Which method does not use the mole ratio in an appropriate manner? Explain.
Toby's method does not correctly use the mole ratio that is required.
- Two of the methods in Model 2 give the same answer. Show that they are mathematically equivalent methods.

$$\begin{aligned} &\text{Jerry's method} \\ &\left\{ \begin{aligned} &50.0 \times \frac{1}{17.0} \times \frac{1}{2} \times 28.0 = 21.0 \text{ g } \text{N}_2 \\ &\quad \quad \quad \downarrow \\ &\quad \quad \quad 1.5 \times \frac{1}{2} \\ &\quad \quad \quad \downarrow \\ &0.75 \times 28.0 = 21.0 \text{ g } \text{N}_2 \end{aligned} \right. \end{aligned}$$

- Use either Rachel or Jerry's method from Model 2 to calculate the mass of hydrogen needed to make 50.0 g of ammonia. $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$

$$50.0 \text{ g} \left(\frac{1 \text{ mole}}{17.0 \text{ g}} \right) \left(\frac{3 \text{ moles}}{2 \text{ moles}} \right) \left(\frac{2.0 \text{ g}}{1 \text{ mole}} \right) = 21.0 \text{ g } \text{H}_2 \text{ required}$$

Pogil Mole Ratios Model 2

Lifelique



Pogil Mole Ratios Model 2:

Mole Ratios Lifeliqe, 2019 This lesson plan relates balanced chemical equations to everyday analogies defines stoichiometry and uses mole ratios to convert between amounts of substances in chemical reactions

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