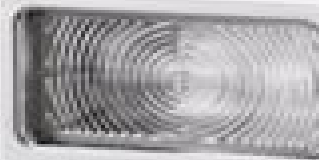


 **Polaroid**

**a530**  
5.0 MEGAPIXELS

...



# Polaroid Camera A530 Manual

**Lou Jacobs**



## **Polaroid Camera A530 Manual:**

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web limiting reactant theoretical and percentage yields key  $\text{K}_2\text{S} + 2\text{H}_2\text{O} \rightarrow 4\text{KOH} + 3\text{O}_2$  if a reaction vessel contains 0 mol  $\text{K}_2\text{O}$  and 0 mol  $\text{H}_2\text{O}$  what is the limiting reactant

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 $\text{Cu}_2\text{S} + 3\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu}_2\text{S} + 3\text{AgNO}_3$

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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$  problems 6.8 are similar to problem 5

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web limiting reagents problem one 1 for the balanced equation  $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$  if the reaction of 21.3 grams of  $\text{O}_2$  produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have  $\text{O}_2$  16.232 grams mole 21.3 grams 0.66 moles 32 g/mol

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0 g  $\text{Al}$   $\times \frac{1 \text{ mol Al}}{2 \text{ mol Al}_2\text{O}_3} \times 101.964 \text{ g Al}_2\text{O}_3 = 302.3 \text{ g Al}_2\text{O}_3$  c determine the percent yield yield actual  $\times 100$  260.0 g  $\times 100$  86.01 theoretical 302.3 g

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web 7 Na<sub>2</sub>C<sub>2</sub>O<sub>4</sub> is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H<sub>2</sub> are left over 13

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