

Unit 7 Quiz Study Guide – Lewis Structures, Polarity & VSEPR

Draw the lewis dot diagram for the following elements:

1) Na



2) Cl



3) He



4) P



5) Se



6) What are the steps for drawing a Lewis Structure?

1. **Count up valence electrons – SHOW WORK**
2. **Determine Central Atom – C, singular/alone, more electronegativity**
3. **Bond terminal atoms**
4. **Recount**
5. **Fill in lone pairs outside first**
6. **Check if happy (STABLE OCTET. Only H, He, and B don't need 8. Everyone else NEEDS 8.)**

7) Define polarity: **unequal sharing of electrons in a polar bond. Polar molecules have an overall dipole. They are not symmetrical and they "move" (horse and cart analogy)**

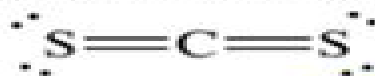
8) Define electronegativity: **Atoms desire to get electrons. Fluorine is the most electronegative.**

For each of the following pairs of molecules, draw the lewis structure. Then, determine which is most polar and explain your reason for making this choice:

9) carbon disulfide (CS_2)

OR

sulfur difluoride (SF_2)

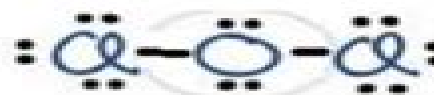


Both are nonpolar!

10) nitrogen trichloride (NCl_3)

OR

oxygen dichloride (OCl_2)



Nitrogen trichloride because it is not symmetrical and "would move"

Polarity Study Guide

N Noddings



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