

3.48 Strategy: Tin(II) fluoride is composed of Sn and F. The mass due to F is based on its percentage by mass in the compound. How do we calculate mass percent of an element?

Solution: First, we must find the mass % of fluorine in SnF_2 . Then, we convert this percentage to a fraction and multiply by the mass of the compound (24.6 g), to find the mass of fluorine in 24.6 g of SnF_2 .

The percent by mass of fluorine in tin(II) fluoride, is calculated as follows:

$$\begin{aligned}\text{mass \% F} &= \frac{\text{mass of F in 1 mol SnF}_2}{\text{molar mass of SnF}_2} \times 100\% \\ &= \frac{2(19.00 \text{ g})}{156.7 \text{ g}} \times 100\% = 24.23\% \text{ F}\end{aligned}$$

Converting this percentage to a fraction, we obtain $24.23/100 = 0.2423$.

Next, multiply the fraction by the total mass of the compound.

$$7 \text{ g F in } 24.6 \text{ g SnF}_2 = (0.2423)(24.6 \text{ g}) = 5.97 \text{ g F}$$

Check: As a ball-park estimate, note that the mass percent of F is roughly 25 percent, so that a quarter of the mass should be F. One quarter of approximately 24 g is 6 g, which is close to the answer.

Note: This problem could have been worked in a manner similar to Problem 3.46. You could complete the following conversions:

$$\text{g of SnF}_2 \rightarrow \text{mol of SnF}_2 \rightarrow \text{mol of F} \rightarrow \text{g of F}$$

3.49 In each case, assume 100 g of compound.

$$(a) \quad 2.1 \text{ g H} \times \frac{1 \text{ mol H}}{1.008 \text{ g H}} = 2.08 \text{ mol H}$$

$$65.3 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 4.081 \text{ mol O}$$

$$32.6 \text{ g S} \times \frac{1 \text{ mol S}}{32.07 \text{ g S}} = 1.017 \text{ mol S}$$

This gives the formula $\text{H}_{2.08}\text{O}_{4.081}\text{S}_{1.017}$. Dividing by 1.017 gives the empirical formula, HS_2O_4 .

$$(b) \quad 26.2 \text{ g Al} \times \frac{1 \text{ mol Al}}{26.98 \text{ g Al}} = 0.9707 \text{ mol Al}$$

$$79.8 \text{ g Cl} \times \frac{1 \text{ mol Cl}}{35.45 \text{ g Cl}} = 2.251 \text{ mol Cl}$$

This gives the formula, $\text{Al}_{0.9707}\text{Cl}_{2.251}$. Dividing by 0.9707 gives the empirical formula, AlCl_2 .

3.50 (a)

Strategy: In a chemical formula, the subscripts represent the ratio of the number of moles of each element that combine to form the compound. Therefore, we need to convert from mass percent to moles in order to determine the empirical formula. If we assume an exactly 100-g sample of the compound, do we know the mass of each element in the compound? How do we then convert from grams to moles?

Solution: If we have 100 g of the compound, then each percentage can be converted directly to grams. In this sample, there will be 40.3 g of C, 6.6 g of H, and 53.3 g of O. Because the subscripts in the formula

Raymond Chang Solution Manual

BM King



Raymond Chang Solution Manual:

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