

CHAPTER 10: Reactions

Reaction Energy**SECTION 1**

Write a balanced chemical equation for the following reactions in the space provided.

- For elements in their standard state, the value of ΔH_f° is _____.
- The formation and decomposition of a substance is represented by the following thermodynamic equations:



- Is the energy being added to or is it being subtracted from the decomposition?
- What is the correct sign for the enthalpy change in this decomposition reaction?

Write the correct answer in the space to the left. Show all your work in the space provided.

- _____ $\Delta H_{\text{f}}^\circ$ of a reaction is 2074°C whereas a 4.000 g of substance reacts with 10.00 g of O_2 to form 14.00 g of decomposition product.
- _____ A substance has a specific heat of $0.001 \text{ J/g}^\circ\text{C}$. If a 1.0 g sample undergoes a 1.0°C increase in temperature, it requires 1.0 J of energy. How much heat is required to raise the temperature of 1.0 g of substance from 20°C to 100°C ?
- The product is a molecular form of substance, and the substance has an enthalpy of 1000 kJ/mol .
 - What is the value of ΔH for this reaction?

Reaction Energy Section 1 Answer Key

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Gallagher, Paul Ingram**



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