Worksheet 1-1 Measuring Reaction Rates

Due Date -

Answer the questions in the space provided. You must show all of your work to receive full marks. All answers MUST be rounded to the correct number of Sig. Figures.

A chemist wishes to determine the rate of reaction of zinc with hydrochloric acid. The equation for the reaction is:

$$Zn_{(s)} + 2HCl_{(aq)} \rightarrow H_{2(g)} + ZnCl_{2(aq)}$$

A piece of zinc is dropped into 1.00 L of 0.100 M HCl and the following data were obtained:

Time	Mass of Zinc
0 s	0.016 g
4 s	0.014 g
8 s	0.012 g
12 s	0.010 g
16 s	0.008 g
20 s	0.006 g

-0.0065

a) Calculate the Rate of Reaction in grams of Zn consumed per second.

Answer 5.0×10-4

b) Calculate the Rate of Reaction in moles of Zn consumed per second.

Answer 7.6 × 10 Sec

c) Write out the complete ionic equation for the reaction.

- e) What will happen to the [CI] as the reaction proceeds? __nothing , it is a spectator

Reaction Rates Practice Problems Answers

Rob Ritchie

Reaction Rates Practice Problems Answers:

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