



Nonlinear States and Equilibrium Test Review

Directions: Please answer the following questions to the best of your ability.

- Which three laws govern the state of the following?
 - Equilibrium being that which is in state and that forward reaction is spontaneous that rate of the reverse reaction.
 - Three factors that can increase the rate of a reaction are increasing temperature, concentration, and pressure.
 - The activation energy is the initial amount of energy needed for a reaction to take place.
 - The difference in energy between the reactants and products is called the enthalpy change.
 - Equilibrium is the state where the forward and reverse reactions are balanced.
 - An example of a system in equilibrium is a mixture of water and ice. The system is in equilibrium with the surroundings that have the same temperature and pressure.
 - If a reaction is exothermic, the forward reaction, which is exothermic, will be favored with more heat. The reaction will shift to the right.
- For the reaction, fill in the blank (L, P, R) for the following changes:

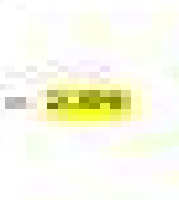
$$\text{N}_2\text{O}_4 \rightleftharpoons 2\text{NO}_2$$

a) <u>R</u> increase in pressure	b) <u>L</u> decrease in P
c) <u>R</u> decrease in temperature	d) <u>L</u> addition
e) <u>R</u> addition of a catalyst	f) <u>R</u> increase NO_2
g) <u>L</u> increase N ₂	h) <u>R</u> decrease pressure
- In a 1.0 L container at equilibrium, 0.20 moles of N_2O_4 , 0.10 moles NO_2 , 0.10 moles CO_2 , and 0.10 moles H_2 . Calculate the value of the equilibrium constant for the reaction.

$$\text{N}_2\text{O}_4 \rightleftharpoons 2\text{NO}_2$$

$\text{N}_2\text{O}_4 = 0.20 \text{ moles} / 1.0 \text{ L} = 0.20 \text{ M}$
 $\text{NO}_2 = 0.10 \text{ moles} / 1.0 \text{ L} = 0.10 \text{ M}$
 $\text{CO}_2 = 0.10 \text{ moles} / 1.0 \text{ L} = 0.10 \text{ M}$
 $\text{H}_2 = 0.10 \text{ moles} / 1.0 \text{ L} = 0.10 \text{ M}$

$$K_c = \frac{[\text{NO}_2]^2}{[\text{N}_2\text{O}_4]} = \frac{(0.10)^2}{(0.20)} = 0.05$$



Reaction Rates And Equilibrium Test And Answers

Neil D. Jespersen, Pamela Kerrigan



Reaction Rates And Equilibrium Test And Answers:

Solutions of Reaction Rate Equations Pertaining to Electron Irradiation of 4:1 Mixtures of N₂ and O₂

Frankline E. Niles, Edna L. Lortie, 1967 One of the Keneshea computer codes see AD 424 173 was adapted for use on the Ballistic Research Laboratories Electronic Scientific Computer Using this modified code reaction rate equations were solved for the following 15 species e NO₂ O O₂ O₃ N₂ NO O₂ N NO N₂O NO₂ O and O₃ The calculations were made for a 4 : 1 mixture of N₂ and O₂ at 1 torr total pressure and 300K Rate constants as given by Keneshea and Fowler see AD 646 975 were used The solutions are presented as number densities versus time after the start of the irradiating electron beam A description of the modified code is presented Author [A Level Chemistry MCQ \(Multiple Choice Questions\)](#) Arshad Iqbal, 2019-06-18 The A Level Chemistry Multiple Choice Questions MCQ Quiz with Answers PDF A Level Chemistry MCQ PDF Download Quiz Questions Chapter 1 28 Practice Tests with Answer Key IGCSE GCE Chemistry Questions Bank MCQs Notes includes revision guide for problem solving with hundreds of solved MCQs A Level Chemistry MCQ with Answers PDF book covers basic concepts analytical and practical assessment tests A Level Chemistry MCQ PDF book helps to practice test questions from exam prep notes The A Level Chemistry MCQs with Answers PDF eBook includes revision guide with verbal quantitative and analytical past papers solved MCQs A Level Chemistry Multiple Choice Questions and Answers MCQs PDF Free download chapter 1 a book covers solved quiz questions and answers on chapters Alcohols and esters atomic structure and theory benzene chemical compound carbonyl compounds carboxylic acids acyl compounds chemical bonding chemistry of life electrode potential electrons in atoms enthalpy change equilibrium group IV groups II and VII halogenoalkanes hydrocarbons introduction to organic chemistry ionic equilibria lattice energy moles and equations nitrogen and sulfur organic and nitrogen compounds periodicity polymerization rates of reaction reaction kinetics redox reactions and electrolysis states of matter transition elements tests for college and university revision guide A Level Chemistry Quiz Questions and Answers PDF free download eBook s sample covers beginner s solved questions textbook s study notes to practice online tests The book IGCSE GCE Chemistry MCQs Chapter 1 28 PDF includes high school question papers to review practice tests for exams A Level Chemistry Multiple Choice Questions MCQ with Answers PDF digital edition eBook a study guide with textbook chapters tests for IGCSE NEET MCAT GRE GMAT SAT ACT competitive exam A Level Chemistry Mock Tests Chapter 1 28 eBook covers problem solving exam tests from chemistry textbook and practical eBook chapter wise as Chapter 1 Alcohols and Esters MCQ Chapter 2 Atomic Structure and Theory MCQ Chapter 3 Benzene Chemical Compound MCQ Chapter 4 Carbonyl Compounds MCQ Chapter 5 Carboxylic Acids and Acyl Compounds MCQ Chapter 6 Chemical Bonding MCQ Chapter 7 Chemistry of Life MCQ Chapter 8 Electrode Potential MCQ Chapter 9 Electrons in Atoms MCQ Chapter 10 Enthalpy Change MCQ Chapter 11 Equilibrium MCQ Chapter 12 Group IV MCQ Chapter 13 Groups II and VII MCQ Chapter 14 Halogenoalkanes MCQ Chapter 15 Hydrocarbons MCQ Chapter 16 Introduction to Organic Chemistry MCQ

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