



Honors/Advanced Placement Equilibrium Test Review

Directions: Please answer the following questions to the best of your ability.

1. Which three laws govern the aspects of the following?
 - a) Equilibrium: that which is stable (has not yet that forward reaction to the other side of the reaction equation).
 - b) Three factors that can increase the value of a reaction are increasing _____ and _____.
 - c) The activation energy is the _____ amount of energy needed for reactants to become products.
 - d) The difference in energy between the reactants and products is called the _____.
 - e) Activation energy for the chemical reaction is the _____.
 - f) An increase in pressure for a system at equilibrium will cause the reaction to shift in the direction that has _____ number of gaseous molecules.
 - g) If a reaction is endothermic in that forward direction, when an increase in temperature will cause the reaction to shift to the _____.
2. For the reaction, fill in the way it will shift. (L, R, N) for the following changes:

$$\text{N}_2\text{g} + 3\text{H}_2\text{g} \rightleftharpoons 2\text{NH}_3\text{g} + 92 \text{ kJ}$$

a) _____ increase in pressure b) _____ decrease in temperature c) _____ addition of a catalyst d) _____ increase both	e) _____ decrease N₂ f) _____ addition g) _____ increase N₂, H₂ h) _____ decrease pressure
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3. In a 5.0 L container at equilibrium contains 1.00 moles of N₂, 0.50 moles H₂, 0.50 moles NH₃, and 0.50 moles H₂. Calculate the value of the equilibrium constant for this reaction.

$$\text{N}_2\text{g} + 3\text{H}_2\text{g} \rightleftharpoons 2\text{NH}_3\text{g}$$
4. The following reaction has an equilibrium constant of 1.1. What is the concentration of N₂, if [CO₂] = 0.10 M, [CO] = 1.70 M, and [O₂] = 1.00 M.

$$2 \text{COg} + \text{O}_2\text{g} \rightleftharpoons 2\text{CO}_2\text{g}$$

Reaction Rates And Equilibrium Test A Answers

Patrick Vollmar



Reaction Rates And Equilibrium Test A Answers:

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