

Name: \_\_\_\_\_ per: \_\_\_\_\_

### Worksheet - Reaction Rates

Use this reaction for the questions below:  $C_6H_{12}O_6 (s) + 6 O_2(g) \rightarrow 6 H_2O (g) + 6 CO_2 (g)$

- What happens to the concentrations of:
  - $C_6H_{12}O_6$  &  $O_2$  as the reaction proceeds  $\rightarrow$ ?
  - $H_2O$  +  $CO_2$  as the reaction proceeds  $\rightarrow$ ?
- According to the collision theory, what 3 circumstances are needed for  $C_6H_{12}O_6$  &  $O_2$  to react?
- What is the activation energy for a chemical reaction?
- Use the equation & the collision theory to explain:



| Change in condition:                                                    | Does this <u>increase</u> or <u>decrease</u> the rate of reaction? | Explain why                                                   |
|-------------------------------------------------------------------------|--------------------------------------------------------------------|---------------------------------------------------------------|
| a. <u>Increasing the temperature</u>                                    | Ex: Increase (springs up)                                          | Ex: Molecules move faster & collide more = increased rxn rate |
| b. <u>Increasing the concentration</u> of $C_6H_{12}O_6$                |                                                                    |                                                               |
| c. <u>Decreasing the concentration</u> of $O_2$                         |                                                                    |                                                               |
| d. <u>Increase the surface area</u> by chewing up food in your mouth.   |                                                                    |                                                               |
| e. <u>Decreasing the temperature</u>                                    |                                                                    |                                                               |
| f. <u>Increasing the pressure</u> in the container                      |                                                                    |                                                               |
| g. <u>Decreasing the concentration</u> of $H_2O$                        |                                                                    |                                                               |
| h. <u>Increasing the volume</u> of the container the reaction occurs in |                                                                    |                                                               |
| i. <u>Increasing the concentration</u> of $CO_2$                        |                                                                    |                                                               |
| j. <u>Using a catalyst</u> (like salivary amylase)                      |                                                                    |                                                               |

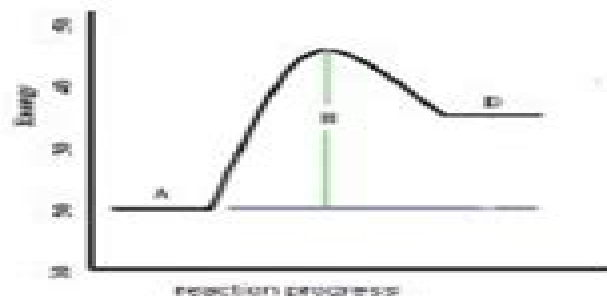
- On the accompanying energy diagram, label the following terms:

a. reactants    b. products    c. activation energy

- On the graph to the right, draw and label what this diagram would look like if a catalyst was added to the reaction.

- Graph reading

- How much energy (kJ) do the reactants have?
- How much energy (kJ) do the products have?
- How much energy (kJ) is required to activate this un-catalyzed reaction?
- Is this reaction endothermic or exothermic? How do you know?



# Reaction Rates And Equilibrium Worksheet Answers

**Clifford Othen**



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*Workbook 15* Thobeka Pete, Turret Correspondence College (Johannesburg), 1988

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**The Equilibrium Assumption in the Theory of Absolute Reaction Rates** Charles Francis Curtiss, 1948 The assumption that systems in the activated state are in equilibrium with the reactant molecules has been open to question In this paper the methods of classical mechanics are used to obtain an expression for the rate constant without making this assumption The result is similar in form to that obtained by Eyring except that the average velocity of the points in configuration space is

related directly to the distribution in velocities of the molecules of the gas in real space If one assumes that this distribution is Maxwellian one obtains the Eyring expression However the chemical reaction itself acts as a perturbing influence and causes the distribution in velocities to be somewhat different from Maxwellian It is this perturbation of the distribution function which leads to a different value of the rate constant and consequently to a measure of the effect of the equilibrium assumption

Author **Rates of Reaction and Equilibria** Clifford Othen,1972 **Theories of Chemical Reaction Rates** Keith James Laidler,1979 2003, [PDF](#) [On the Theory of Chemical Reaction Rates](#) Raphael D. Levine,1964 **Metal-water Reactions** Leo F. Epstein,1959 **Rates of Reaction & Equilibrium** Roy Brown,Irish science teachers' association,1983 *Reaction Rates* ,200? [CHEMICAL KINETICS](#) NARAYAN CHANGDER,2024-05-16 If you need a free PDF practice set of this book for your studies feel free to reach out to me at cbsenet4u gmail com and I ll send you a copy

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