

Reactions in Aqueous Solution

1 2 3 4 5 6 7 8 9 10 11 12 13 14 15 16 17 18 19 20

1. The products of the following reaction, which is the net ionic equation, are:

☐ AgCl
☐ AgNO_3
☒ AgCl and NaNO_3
☐ NaCl and AgNO_3
2. In a double replacement reaction, evidence is observed that the following product is formed in a double replacement:

☐ BaCl_2
☐ H_2O
☒ BaCl_2
☐ AgCl
3. Evidence of the following is observed:

☐ all water containing H^+ ions is acidic
 ☐ all water containing H^+ ions is acidic
 ☒ all strong acids are acidic
 ☐ all strong acids are acidic
4. Write the following chemical equation for the reaction of sodium metal with water. Then, write the net ionic equation.

$$2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$$
5. The balanced equation for the above reaction is:

$$2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$$
6. When many metals of the transition series are heated, they form oxides. Which of the following is the correct formula for the oxide of iron?

☐ Fe_2O_3
☐ Fe_2O_4
☒ Fe_2O_3
☐ Fe_2O_4
7. Sodium hydroxide and hydrochloric acid react to form sodium chloride and water. Write the products of the reaction.

☐ NaCl and H_2O
☐ NaCl
☒ NaCl
☐ H_2O

Reactions In Aqueous Solutions Practice Test Answers

**Barron's Educational Series, Neil D.
Jespersen, Pamela Kerrigan**



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