

3. Refer to the data table in Question 2.

a. How do the reduced ratios in the last column compare to the coefficients in the reaction shown in Model 1?

The ratios are the same as the coefficients in Model 1: 1:3:2.

b. Use mathematical concepts to explain how your answer in part a is possible.

The coefficients were all multiplied by the same number, so they all ended back to the same thing.

4. Even 538 is a small number of molecules to use in a reaction. Typically chemists use much larger numbers of molecules. (Recall that one mole is equal to 6.02×10^{23} particles.) Consider each situation below as it relates to the reaction in Model 1: $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$.

a. Calculate the amount of reactants consumed and products made.

b. Record the ratio of N_2 to H_2 to NH_3 . Reduce the ratio to the lowest whole number possible.

	N_2 Consumed	H_2 Consumed	NH_3 Produced	Ratio $\text{N}_2:\text{H}_2:\text{NH}_3$
If the reaction occurred 6.02×10^{23} times, how many molecules would be consumed or produced?	6.02×10^{23}	1.81×10^{24}	1.20×10^{24}	1 : 3 : 2
How many <i>moles</i> of each substance would be consumed or produced in the previous situation?	1	3	2	1 : 3 : 2

5. Refer to the data table in Question 4.

a. How do the reduced ratios in the last column compare to the coefficients in the reaction in Model 1?

The reduced ratios are the same as the coefficients.

b. Use mathematical concepts to explain how your answer in part a is possible.

We are still multiplying coefficients by the same number so we just didn't everything by the number to get back to a reduced ratio.

6. The ratio obtained from the coefficients in a balanced chemical equation is called the **mole ratio**.

a. What is the mole ratio for the reaction in Model 1?



b. Explain why this ratio is called the mole ratio?

It is the reduced ratio in moles (all the coefficients = 6.02×10^{23}).

Why Mole Ratios Pogil Answer Key

NARAYAN CHANGDER



Why Mole Ratios Pogil Answer Key:

Chemistry Education in the ICT Age Minu Gupta Bhowon, Sabina Jhaumeer-Laulloo, Henri Li Kam Wah, Ponnadurai Ramasami, 2009-07-21 The 20 International Conference on Chemical Education 20 ICCE which had rd th Chemistry in the ICT Age as the theme was held from 3 to 8 August 2008 at Le Meridien Hotel Pointe aux Piments in Mauritius With more than 200 participants from 40 countries the conference featured 140 oral and 50 poster presentations th Participants of the 20 ICCE were invited to submit full papers and the latter were subjected to peer review The selected accepted papers are collected in this book of proceedings This book of proceedings encloses 39 presentations covering topics ranging from fundamental to applied chemistry such as Arts and Chemistry Education Biochemistry and Biotechnology Chemical Education for Development Chemistry at Secondary Level Chemistry at Tertiary Level Chemistry Teacher Education Chemistry and Society Chemistry Olympiad Context Oriented Chemistry ICT and Chemistry Education Green Chemistry Micro Scale Chemistry Modern Technologies in Chemistry Education Network for Chemistry and Chemical Engineering Education Public Understanding of Chemistry Research in Chemistry Education and Science Education at Elementary Level We would like to thank those who submitted the full papers and the reviewers for their timely help in assessing the papers for publication th We would also like to pay a special tribute to all the sponsors of the 20 ICCE and in particular the Tertiary Education Commission <http://tec.intnet.mu> and the Organisation for the Prohibition of Chemical Weapons <http://www.opcw.org> for kindly agreeing to fund the publication of these proceedings **STOICHIOMETRY** NARAYAN CHANGDER, 2024-04-01 Note Anyone can request the PDF version of this practice set workbook by emailing me at cbsenet4u@gmail.com You can also get full PDF books in quiz format on our youtube channel <https://www.youtube.com/smartquiziz> I will send you a PDF version of this workbook This book has been designed for candidates preparing for various competitive examinations It contains many objective questions specifically designed for different exams Answer keys are provided at the end of each page It will undoubtedly serve as the best preparation material for aspirants This book is an engaging quiz eBook for all and offers something for everyone This book will satisfy the curiosity of most students while also challenging their trivia skills and introducing them to new information Use this invaluable book to test your subject matter expertise Multiple choice exams are a common assessment method that all prospective candidates must be familiar with in today's academic environment Although the majority of students are accustomed to this MCQ format many are not well versed in it To achieve success in MCQ tests quizzes and trivia challenges one requires test taking techniques and skills in addition to subject knowledge It also provides you with the skills and information you need to achieve a good score in challenging tests or competitive examinations Whether you have studied the subject on your own read for pleasure or completed coursework it will assess your knowledge and prepare you for competitive exams quizzes trivia and more **You Can Do Chemistry** Sunil Tanna, 2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as

well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018

Why Mole Ratios Pogil Answer Key Book Review: Unveiling the Power of Words

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