



Chemical Changes

Read the passage below and answer the following questions:



1. What are the products of this chemical reaction?

Mercury and oxygen

2. What is the reactant? Mercury oxide

3. What is the physical state of the reactant and the products?

The reactant is a solid, the products are a liquid and a gas.

4. According to the law of conservation of mass, if 100g of mercury oxide is heated, how much mercury and oxygen will be produced?

100g



5. What products are formed? Mercury and oxygen

6. What mass of mercury is formed? 80g

7. What are the physical states of the products? Mercury is a liquid, oxygen is a gas.

8. What is the physical state of the reactant? Mercury oxide is a solid.

9. What is the mass of the products? 100g

10. How many moles of the products are formed? 1 mole

Read the passage below and answer the following questions:

11. How many moles of H_2O are produced if 1 mole of H_2 reacts with 1 mole of O_2 ?

2 moles of H_2O are produced.

Skills Directed A Section Chemical Changes Answers

Shasha Hu



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