

Problem #2: It takes 180mL of 1.2 M HCl to neutralize a 0.75 M solution of NaOH. How much NaOH did I start with?

$$V_a = 180 \text{ mL}$$

$$M_a = 1.2 \text{ M}$$

$$V_b = ?$$

$$M_b = 0.75$$

$$V_a \times M_a = V_b \times M_b$$

$$180 \text{ mL} \times 1.2 \text{ M} = V_b \times 0.75 \text{ M}$$

$$\frac{216}{0.75 \text{ M}} = \frac{V_b \times 0.75 \text{ M}}{0.75 \text{ M}}$$

$$288 \text{ mL}$$

Titration Problems With Answers

**Robert J. Salm, Robert J. Salm Robert J.
Salm Ph.D.**

Titration Problems With Answers:

Chemical Calculations with Explanatory Notes, Problems, and Answers, Specially Adapted for Use in Colleges and Science Schools Richard Lloyd Whiteley, 1896 **Clinical Calculations** Joyce LeFever Kee, Sally M. Marshall, 2016-01-25

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