

On donne la température normale de fusion du phosphore : $t_f = 44^\circ\text{C}$; la capacité thermique massique du phosphore liquide : $c_l = 0,25 \text{ cal}\cdot^\circ\text{C}^{-1}\cdot\text{g}^{-1}$; la capacité thermique massique du phosphore solide : $c_s = 0,20 \text{ cal}\cdot^\circ\text{C}^{-1}\cdot\text{g}^{-1}$; la chaleur latente massique de fusion du phosphore à 44°C : $L_{fs} = 4,7 \text{ cal}\cdot\text{g}^{-1}$.
Calculer la chaleur latente massique de fusion du phosphore à la température t du phosphore surfondu : L_{fs} en fonction de la température t et des données.

Exercice 8.

Un calorimètre de capacité thermique totale : C (y compris celle des corps qui y sont contenus) dont la température initiale est θ_0 se refroidit par perte thermique. La puissance de fuite est proportionnelle à l'écart entre la température θ du calorimètre à chaque instant et la température extérieure constante θ_a .

1. Etablir la loi de variation θ en fonction du temps t , de C , de a coefficient de proportionnalité de la puissance de fuite, et des températures θ_0 et θ_a .
2. A.N. : $C = 200 \text{ cal}\cdot^\circ\text{C}^{-1}$; $\theta_0 = 20^\circ\text{C}$; $\theta_a = 10^\circ\text{C}$; au bout d'une minute la température est de $19,5^\circ\text{C}$; calculer le temps au bout duquel elle sera de 15°C .

Exercice 9.

On considère un calorimètre et l'eau contenue, de capacité thermique totale $C = 400 \text{ cal}\cdot^\circ\text{C}^{-1}$ (on rappelle $1 \text{ cal} = 4,185 \text{ J}$) et de température initiale $\theta_0 = 15^\circ\text{C}$, dans lequel on immerge un serpent, de capacité thermique négligeable, parcouru par un liquide de capacité thermique massique constante $c = 0,4 \text{ cal}\cdot^\circ\text{C}^{-1}\cdot\text{g}^{-1}$, avec un débit régulier $k = 1 \text{ g}\cdot\text{s}^{-1}$. Ce liquide entre à la température $\theta_1 = 80^\circ\text{C}$ constante et sort à la température θ_2 du calorimètre. On négligera les fuites thermiques.

1. Etablir la relation donnant θ_2 en fonction du temps t .
2. Quelle serait la température du calorimètre initialement à θ_0 , si l'on versait directement 100 g du liquide dont la température initiale était θ_1 , le serpent étant alors vide ?
3. Le serpent est maintenant parcouru par de l'hydrogène qui entre à la température $\theta_1 = 80^\circ\text{C}$ et sort à la température θ_2 du calorimètre, dans les mêmes conditions initiales que précédemment ; au bout du temps $t = 100 \text{ s}$, on note $\theta_2 = 52^\circ\text{C}$. Déterminer la capacité thermique spécifique de l'hydrogène, le débit est $k = 1 \text{ g}\cdot\text{s}^{-1}$.

Premier principe.

Exercice 10.

L'état initial d'une mole de gaz parfait est caractérisé par $P_0 = 2 \cdot 10^5 \text{ Pa}$, $V_0 = 14 \text{ L}$. On fait subir successivement à ce gaz les transformations réversibles suivantes :

- une détente isobare qui double son volume ;
 - une compression isotherme qui le ramène à son volume initial ;
 - un refroidissement isochore qui le ramène à l'état initial.
1. A quelle température s'effectue la compression isotherme ? En déduire la pression maximale atteinte.
 2. Représenter le cycle de transformations dans le diagramme (P, V) .
 3. Calculer les travaux et transferts thermiques échangés par le système au cours du cycle, soient W_1 , W_2 , W_3 , Q_1 , Q_2 et Q_3 en fonction de P_0 , V_0 et $\gamma = \frac{C_p}{C_v} = 1,4$ (supposé constant dans le domaine de températures étudié).
 4. Vérifier $U = 0$ pour le cycle.

Exercice 11 : détente quasi-statistique polytropique d'un gaz parfait.

On considère la détente polytropique d'indice γ constant (transformation pour laquelle le volume V et la pression P vérifient $P V^\gamma = \text{cte}$, avec γ constante positive, on pourra à ce sujet se reporter à l'exercice 7 de la série 23) d'un gaz parfait le menant d'un état

(P_1, V_1, T_1) à un état (P_2, V_2, T_2) avec $V_2 > V_1$. On pose $\gamma = \frac{C_p}{C_v}$ (supposé constant),

Pour quelles valeurs du coefficient γ , la détente du gaz s'accompagne-t-elle :

1. d'absorption de chaleur et d'échauffement du gaz ?
2. d'absorption de chaleur et de refroidissement du gaz ?

Exercice 12 : transformation adiabatique.

Un gaz parfait passe d'un état (P_1, V_1, T_1) à un état (P_2, V_2, T_2) suivant une transformation adiabatique.

On pose $\gamma = \frac{C_p}{C_v}$ (supposé constant).

1. Montrer que, s'il existe une suite continue d'états d'équilibre thermodynamiques internes au cours de la transformation, la pression P et le volume V du gaz sont reliés par : $P V^\gamma = \text{cte}$.
 2. Le gaz est comprimé et passe de la pression P_1 à la pression $P_2 = 2 P_1$. Calculer le travail échangé par le gaz et le milieu extérieur en fonction de P_1 , V_1 et γ .
- Données : $P_1 = 1 \text{ bar}$, $V_1 = 1 \text{ dm}^3$ et $\gamma = 1,4$.

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J Rink



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