Thermochemistry Practice Problems



Flow of Energy:

How many kJ of energy are in a donut that contains 200.0 Calories?

2. What is the specific heat of a substance that has a mass of 25.0 g and requires 525.0 calories to raise its temperature by 15.0°C? Q= mc AT

$$m = 25.0g$$

 $q = 525.0 \text{ cal}$ $525.0 \text{ cal} = (25.0 \text{ g} \text{X} \text{ c} \text{X} (15.0 \text{C})$
 $\Delta T = 15.0 \text{C}$ $C = 1.40^{\text{Cal}}/\text{g} \cdot \text{C}$

 Suppose 100.0 g of H₂O(s) absorbs 1255.0 J of heat. What is the corresponding temperature change? The specific heat capacity of $H_2O(s)$ is 2.1 J/g^*C . $Q = mC \Delta T$

4. How many joules of heat energy are required to raise the temperature of 100.0 g of aluminum by 120.0°C? The specific heat capacity of aluminum is 0.90 J/g°C.

5. In reference to #4, what is the molar heat capacity of aluminum?

Measuring and Expressing Heat Changes:

A student mixed 75.0 mL of water containing 0.75 mol HCl at 25°C with 75.0 mL of water containing 0.75 mol of NaOH at 25°C in a foam cup calorimeter. The temperature of the resulting solution increased to 35°C. How much heat in kilojoules was released by this reaction (Couter = 4.18 J/g°C)? 9=mcot

7. Calculate the amount of heat evolved when 150 g of water are reacted with excess calcium according to the following thermochemical equation:

$$Ca(s) + 2H_2O(l) \rightarrow Ca(OH)_2(s) + H_2(g)$$
 $\Delta H = -65.2 \text{ W}$
 $150_3 H_2O \mid ImbiH_2O \mid -65.2 \text{ KJ} = \left[-370 \text{ KJ} \right]$ exothermic $\left[13.02gH_2O \mid 2mol H_2O \right] = \left[-370 \text{ KJ} \right]$

Thermochemistry Practice Problems And Answers

Daniel C. Tofan

Thermochemistry Practice Problems And Answers:

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Thermochemistry Practice Problems And Answers Book Review: Unveiling the Magic of Language

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