

4

Thermodynamics

Contents and Concepts

- 4.1 Introduction
- 4.2 Thermal Equilibrium
- 4.3 Zeroth Law of Thermodynamics
- 4.4 Heat, Internal Energy and Work
- 4.5 First Law of Thermodynamics
- 4.6 Thermodynamic State Variables

- 4.7 Thermodynamic Processes
- 4.8 Heat Engines
- 4.9 Refrigerators and Heat Pumps
- 4.10 Second Law of Thermodynamics
- 4.11 Carnot Cycle and Carnot Engine

4.1 Introduction

Thermodynamics is the study of conversion of work (or any form of energy) into heat and the other way round. It is an important branch of physics having many practical applications.

According to modern theory, heat is a form of energy. When a hot object is in contact with a cold object, the hot object gets cooler and the cold object becomes warmer, and both the objects reach the same temperature after some time due to exchange of heat.

Q.1. What is thermodynamics?

Ans: Thermodynamics is the branch of physics that deals with the concepts of heat and temperature and the inter-conversion of heat and other forms of energy.

4.2 Thermal Equilibrium

Q.2. Can you recall? (Textbook page no. 73)

When a piece of ice is placed in water at room temperature, the ice melts and water cools down. Why does their temperature change?

Ans:

- i. Water at room temperature is at higher temperature than the temperature of ice.
- ii. When ice is placed in the water, the heat energy is transferred from water at higher temperature to ice at lower temperature.
- iii. In this way, the temperature of ice increases and it melts whereas, the temperature of water decreases and it cools down.
- iv. This temperature change takes place due to exchange of heat.

Q.3. When do two systems are said to be in thermal equilibrium?

Ans:

- i. Two systems in thermal contact with each other are said to be in thermal equilibrium if they do not transfer heat between each other.
- ii. When two objects are at the same temperature, they are in thermal equilibrium.

Q.4. Can you tell? (Textbook page no. 76)

Why different objects kept on a table at room temperature do not exchange heat with the table?

Ans:

All the objects kept on table along with the table are at same temperature (room temperature) i.e., all the objects are in thermal equilibrium with the table. Hence, they do not exchange heat with the table.

4.3 Zeroth Law of Thermodynamics

Q.5. State zeroth law of thermodynamics.

Ans: If two systems are each in thermal equilibrium with a third system, they are also in thermal equilibrium with each other.

Q.6. Explain zeroth law of thermodynamics using a schematic representation.

Ans: **Statement:** If two systems are each in thermal equilibrium with a third system, they are also in thermal equilibrium with each other.



Schematic representation of zeroth law of thermodynamics

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 - The Rise of Digital Reading Thermodynamics Target Publications Std 12 Th
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermodynamics Target Publications Std 12 Th
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermodynamics Target Publications Std 12 Th
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermodynamics Target Publications Std 12 Th
 - Personalized Recommendations
 - Thermodynamics Target Publications Std 12 Th User Reviews and Ratings
 - Thermodynamics Target Publications Std 12 Th and Bestseller Lists
5. Accessing Thermodynamics Target Publications Std 12 Th Free and Paid eBooks

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 - Thermodynamics Target Publications Std 12 Th Budget-Friendly Options
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- ePub, PDF, MOBI, and More
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 - Thermodynamics Target Publications Std 12 Th Enhanced eBook Features
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 - Highlighting and Note-Taking Thermodynamics Target Publications Std 12 Th
 - Interactive Elements Thermodynamics Target Publications Std 12 Th
8. Staying Engaged with Thermodynamics Target Publications Std 12 Th
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermodynamics Target Publications Std 12 Th
9. Balancing eBooks and Physical Books Thermodynamics Target Publications Std 12 Th
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermodynamics Target Publications Std 12 Th
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Thermodynamics Target Publications Std 12 Th
- Setting Reading Goals Thermodynamics Target Publications Std 12 Th
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Thermodynamics Target Publications Std 12 Th
- Fact-Checking eBook Content of Thermodynamics Target Publications Std 12 Th
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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