

# Chemistry Spring Final Study Guide

### Unit 10: Acids & Bases

Use the following information and complete the structure of Acids & Bases.

Acids	Bases
<p><b>Unit 10: Acids &amp; Bases</b></p> <p>Draw a line on the graph and indicate its dependence for each property: Temperature, Volume, and Pressure below.</p> <p>Label the y-axis.</p> <p>Label the x-axis.</p> <p>1. A substance of the particles.</p> <p>2. A substance of the particles.</p> <p>3. A substance of the particles.</p> <p>4. The gas is in.</p> <p>5. The volume.</p>	<p>1. A substance of the particles.</p> <p>2. A substance of the particles.</p> <p>3. A substance of the particles.</p> <p>4. The gas is in.</p> <p>5. The volume.</p>

**Unit 10: Acids & Bases**

Write the definition:

- Acid
- Base
- Indicator
- Titration
- Endpoint
- Equivalence
- Titration

Use the equation that:

Write the correct formula:

1. Write the correct formula of the acid.

2. Write the correct formula of the base.

3. Write the correct formula of the salt.

### Unit 11: Chemical Reactions

Write the balanced equation for each reaction.

1.  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

2.  $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

3.  $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$

4.  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$

5.  $\text{Ca} + \text{O}_2 \rightarrow \text{CaO}$

6.  $\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$

7.  $\text{K} + \text{O}_2 \rightarrow \text{K}_2\text{O}$

8.  $\text{Na} + \text{O}_2 \rightarrow \text{Na}_2\text{O}$

9.  $\text{Li} + \text{O}_2 \rightarrow \text{Li}_2\text{O}$

10.  $\text{Ba} + \text{O}_2 \rightarrow \text{BaO}$

11.  $\text{Sr} + \text{O}_2 \rightarrow \text{SrO}$

12.  $\text{Ca} + \text{O}_2 \rightarrow \text{CaO}$

13.  $\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$

14.  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$

15.  $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$

16.  $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

17.  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

18.  $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$

19.  $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$

20.  $\text{P} + \text{O}_2 \rightarrow \text{P}_2\text{O}_5$

21.  $\text{N} + \text{O}_2 \rightarrow \text{NO}_2$

22.  $\text{Cl}_2 + \text{H}_2 \rightarrow \text{HCl}$

23.  $\text{Br}_2 + \text{H}_2 \rightarrow \text{HBr}$

24.  $\text{I}_2 + \text{H}_2 \rightarrow \text{HI}$

25.  $\text{F}_2 + \text{H}_2 \rightarrow \text{HF}$

26.  $\text{H}_2 + \text{Cl}_2 \rightarrow \text{HCl}$

27.  $\text{H}_2 + \text{Br}_2 \rightarrow \text{HBr}$

28.  $\text{H}_2 + \text{I}_2 \rightarrow \text{HI}$

29.  $\text{H}_2 + \text{F}_2 \rightarrow \text{HF}$

30.  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

31.  $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$

32.  $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$

33.  $\text{P} + \text{O}_2 \rightarrow \text{P}_2\text{O}_5$

34.  $\text{N} + \text{O}_2 \rightarrow \text{NO}_2$

35.  $\text{Cl}_2 + \text{H}_2 \rightarrow \text{HCl}$

36.  $\text{Br}_2 + \text{H}_2 \rightarrow \text{HBr}$

37.  $\text{I}_2 + \text{H}_2 \rightarrow \text{HI}$

38.  $\text{F}_2 + \text{H}_2 \rightarrow \text{HF}$

39.  $\text{H}_2 + \text{Cl}_2 \rightarrow \text{HCl}$

40.  $\text{H}_2 + \text{Br}_2 \rightarrow \text{HBr}$

41.  $\text{H}_2 + \text{I}_2 \rightarrow \text{HI}$

42.  $\text{H}_2 + \text{F}_2 \rightarrow \text{HF}$

43.  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

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98.  $\text{P} + \text{O}_2 \rightarrow \text{P}_2\text{O}_5$

99.  $\text{N} + \text{O}_2 \rightarrow \text{NO}_2$

100.  $\text{Cl}_2 + \text{H}_2 \rightarrow \text{HCl}$

101.  $\text{Br}_2 + \text{H}_2 \rightarrow \text{HBr}$

102.  $\text{I}_2 + \text{H}_2 \rightarrow \text{HI}$

103.  $\text{F}_2 + \text{H}_2 \rightarrow \text{HF}$

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**J Rink**



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*Emerging Trends in Environmental Biotechnology*  
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