

- The particles in a gas are more spread out than the particles in a liquid and have less attractive forces acting between them; so, if the same amount of energy is transferred to a gas as to a liquid, the kinetic energy of the particles in a gas will increase more than the kinetic energy of particles in a liquid.
- Convection—movement of the hot water as it expands and rises, then cools and contracts, mixes the hot and cold water.
- Students should defend their answers. A heavy cast iron skillet transfers heat evenly, resists temperature changes, and is slow to heat. A thin stainless steel skillet transfers heat quickly, has a high surface temperature directly over the flame, and cools quickly.

SECTION: USING HEAT

- A damp towel has water that absorbs heat energy and evaporates as the air blows through it. Because the air loses heat energy to the water, the air is cooler.
- No, the refrigerator releases heat from inside the box out the back of the unit. The refrigerator is in the room, so the heat is still released in the kitchen.
 - Yes, because shivering causes muscle movements that generate heat.
 - Yes, because the large ears of a jackrabbit act as radiators to radiate heat from the rabbit's blood.
- The two objects that will be at different temperatures, such as an inside wall and an outside wall, should be separated by an air-filled space so that they are not touching, in order to eliminate heat loss by conduction.
- A diesel engine does not use spark plugs. Instead, it compresses the fuel-air mixture so much that a spark is ignited.
- $$T_F = 1.8t + 32.0 = [1.8(-235)] + 32.0 = -391^\circ\text{F}$$

$$T = t + 273 = -235 + 273 = 38\text{ K}$$
- $$T_F = 1.8t + 32.0 = (1.8 \times 15) + 32.0 = 59^\circ\text{F}$$

$$T = t + 273 = 15 + 273 = 288\text{ K}$$
- $$T_F = 1.8t + 32.0 = (1.8 \times 190) + 32.0$$

$$T = t + 273 = 190 + 273 = 463\text{ K}$$

$$= 370^\circ\text{F}$$
- $$t = \frac{(T_F - 32.0)}{1.8} = \frac{(125 - 32.0)}{1.8} = 52^\circ\text{C}$$

$$T = t + 273 = 52 + 273 = 325\text{ K}$$
- $$t = \frac{(T_F - 32.0)}{1.8} = \frac{(17 - 32.0)}{1.8} = -8.3^\circ\text{C}$$

$$T = t + 273 = 265\text{ K}$$

$$t = \frac{(T_F - 32.0)}{1.8} = \frac{(83 - 32.0)}{1.8} = 28^\circ\text{C}$$

$$T = t + 273 = 28 + 273 = 301\text{ K}$$
- $$t = \frac{(T_F - 32.0)}{1.8} = \frac{(-83.0 - 32.0)}{1.8}$$

$$= -64.4^\circ\text{C}$$

$$T = t + 273 = -64.4 + 273 = 209\text{ K}$$
- $$t = \frac{(T_F - 32.0)}{1.8} = \frac{(276 - 32.0)}{1.8}$$

$$= 136^\circ\text{C}$$

$$T = t + 273 = 136 + 273 = 409\text{ K}$$
- $$t = T - 273 = 5100 - 273 = 4800^\circ\text{C}$$

$$T_F = 1.8t + 32.0 = 1.8(4800) + 32.0 = 8700^\circ\text{F}$$
- $$t = T - 273 = 1088 - 273 = 815^\circ\text{C}$$

$$T_F = 1.8t + 32.0 = 1.8(815) + 32.0 = 1.50 \times 10^3^\circ\text{F}$$
- $$t = T - 273 = 3500 - 273 = 3200^\circ\text{C}$$

$$T_F = 1.8t + 32.0 = 1.8(3200) + 32.0 = 5800^\circ\text{F}$$
- $$t = T - 273 = 240 - 273 = -30^\circ\text{C}$$

$$T_F = 1.8t + 32.0 = 1.8(-30) + 32.0 = -20^\circ\text{F}$$
- $$T_F = 1.8t + 32.0 = 1.8(29) + 32.0 = 84^\circ\text{F}$$

$$T_F = 1.8t + 32.0 = 1.8(-11) + 32.0 = 12^\circ\text{F}$$
- $$t = T - 273 = 1.0 \times 10^5 - 273 = 1.0 \times 10^5\text{C}$$
- $$t = T - 273 = 4.25 - 273 = -269^\circ\text{C}$$

$$T_F = 1.8t + 32.0 = 1.8(-269) + 32.0 = -482^\circ\text{F}$$
- $$\Delta t = \frac{(\Delta T_F)}{1.8} = \frac{(49^\circ\text{F})}{1.8} = 27^\circ\text{C}$$
 - $$\text{initial } t = \text{final } t - \Delta t$$

$$= 45^\circ\text{F} - 49^\circ\text{F} = -4^\circ\text{F}$$

Math Skills

TEMPERATURE CONVERSION

- $$T_F = 1.8t + 32.0 = (1.8 \times 453) + 32.0 = 847^\circ\text{F}$$

$$T = t + 273 = 453 + 273 = 726\text{ K}$$

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G Orfield



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Understanding Chemistry Fred M. Dewey,1994

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