# Worksheet #10: Stoichiometry Supplement Limiting Reactant, Theoretical Yield, and Percent Yield Step-by-Step

Consider the reaction between magnesium and nitrogen, which produces magnesium nitride:

$$3Mg(s) + N_2(g) \rightarrow Mg_3N_2(s)$$

Molar Masses needed

Track calculations as needed (if you want)

> Problem A: Calculate Reactant Needed and

	Mg	N <sub>2</sub>	Mg <sub>3</sub> N <sub>2</sub>
g/mol	24.31	28.02	100.95
moles			
grams	14.32		

# **Supplemental Stoichiometry Problems**

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