

**QUESTION 2**

- 2.1 The graph of molecular size versus the boiling point is given below.  
The letters **A**, **B** and **C** represent the compounds  $\text{CH}_4$ ,  $\text{C}_2\text{H}_6$  and  $\text{C}_3\text{H}_8$  respectively.



- 2.1.1 Define the term *boiling point*. (2)
- 2.1.2. Describe the trend in the boiling points of the compounds as shown by the graph. (2)
- 2.1.3. Explain the answer to QUESTION 2.1.2 by referring to MOLECULAR SIZE, TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)
- 2.1.4 Which ONE of the compounds (**A**, **B** or **C**) has the HIGHEST vapour pressure? Explain the answer by referring to the data on the graph. (2)
- 2.2 Consider the two molecules in the table below.

NAME OF SUBSTANCE	FORMULA	MOLACULAR MASS (g)	BOILING POINT ( $^{\circ}\text{C}$ )
Ammonia	$\text{NH}_3$	17	- 33
Phosphine	$\text{PH}_3$	x	- 87,4

- 2.2.1 Determine the molecular mass (x) of phosphine (1)
- 2.2.2 Explain the difference in the boiling points by referring to the TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)

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## **Table of Contents Physical Science Question Papers**

1. Understanding the eBook Physical Science Question Papers
  - The Rise of Digital Reading Physical Science Question Papers
  - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Question Papers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Science Question Papers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Question Papers
  - Personalized Recommendations
  - Physical Science Question Papers User Reviews and Ratings
  - Physical Science Question Papers and Bestseller Lists
5. Accessing Physical Science Question Papers Free and Paid eBooks
  - Physical Science Question Papers Public Domain eBooks
  - Physical Science Question Papers eBook Subscription Services
  - Physical Science Question Papers Budget-Friendly Options
6. Navigating Physical Science Question Papers eBook Formats

- ePub, PDF, MOBI, and More
  - Physical Science Question Papers Compatibility with Devices
  - Physical Science Question Papers Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Physical Science Question Papers
    - Highlighting and Note-Taking Physical Science Question Papers
    - Interactive Elements Physical Science Question Papers
  8. Staying Engaged with Physical Science Question Papers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Physical Science Question Papers
  9. Balancing eBooks and Physical Books Physical Science Question Papers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Physical Science Question Papers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Physical Science Question Papers
    - Setting Reading Goals Physical Science Question Papers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Physical Science Question Papers
    - Fact-Checking eBook Content of Physical Science Question Papers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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