

Chemical Bonding

GROUP STUDY GUIDE

1. Bonding & Stability

Why do atoms bond?

Atoms bond to achieve a more stable electronic configuration, typically by filling their outermost shell (valence shell) with electrons.

Types of Bonding:

- Covalent Bonding:** Atoms share pairs of electrons to form a bond. This results in a molecule where the shared electrons are attracted to both nuclei.
- Ionic Bonding:** One atom transfers electrons to another, resulting in oppositely charged ions that are attracted to each other.
- Metallic Bonding:** Metal atoms release their valence electrons into a 'sea of delocalized electrons' that hold the metal cations together.

Diagram 1: Covalent Bonding

Diagram 1 shows two atoms (represented by red spheres) sharing a pair of electrons (represented by a blue sphere) to form a covalent bond. The resulting molecule is shown with the shared electrons between the two atoms.

Diagram 2: Ionic Bonding

Diagram 2 shows an atom losing an electron to become a cation (positive ion) and another atom gaining an electron to become an anion (negative ion). The resulting ionic bond is shown as the electrostatic attraction between these oppositely charged ions.

Diagram 3: Metallic Bonding

Diagram 3 shows a lattice of metal cations (positive ions) surrounded by a 'sea of delocalized electrons' (represented by a blue cloud). The metallic bond is the attraction between the cations and the electron sea.

2. Molecular Structure

Valence Electrons: Electrons in the outermost shell of an atom that are available for bonding.

Valency: The combining capacity of an element, determined by the number of valence electrons it can lose, gain, or share.

Diagram 4: Lewis Dot Diagrams

Diagram 4 shows Lewis dot diagrams for H_2 , O_2 , and H_2O . Each dot represents a valence electron. Pairs of dots between atoms represent shared electron pairs (covalent bonds).

Diagram 5: VSEPR Theory

Diagram 5 illustrates the Valence Shell Electron Pair Repulsion (VSEPR) theory. It shows how the repulsion between electron pairs (bonding and non-bonding) determines the molecular geometry. Examples include linear (180°), trigonal planar (120°), tetrahedral (109.5°), and bent shapes.

Diagram 6: Molecular Polarity

Diagram 6 shows the difference between polar and non-polar molecules. Polar molecules have an uneven distribution of electron density, leading to partial positive (δ^+) and partial negative (δ^-) charges. Non-polar molecules have a symmetrical distribution of electron density.

3. Bond Properties

Strength of Bond: The energy required to break a bond. Stronger bonds are shorter and have higher bond dissociation energies.

Diagram 7: Bond Length and Energy

Diagram 7 shows a graph of potential energy versus bond length. The minimum energy point corresponds to the equilibrium bond length. Stronger bonds have lower minimum energy and shorter bond lengths.

Diagram 8: Bond Angle

Diagram 8 shows how the repulsion between electron pairs affects the bond angle in a molecule. For example, in water (H_2O), the lone pairs on oxygen repel the bonding pairs, resulting in a bond angle of approximately 104.5° .

4. Chemical Reactions

Reaction Mechanism: A series of steps that describe how a chemical reaction occurs at the molecular level.

Diagram 9: Reaction Mechanism

Diagram 9 shows a reaction mechanism for the reaction between hydrogen peroxide and iodide ions. It includes the reactants, intermediate species, and the final products, with arrows indicating the movement of electron pairs.

5. Summary

Chemical bonding is the process by which atoms combine to form molecules or crystals. The type of bond (covalent, ionic, metallic) depends on the nature of the atoms and the arrangement of their valence electrons. Understanding bonding helps explain the properties of substances, such as their melting points, boiling points, and electrical conductivity.

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Establishment of a National Science Academy United States. Congress. House. Committee on Science and Astronautics, United States. Congress. House. Committee on Science and Astronautics. Subcommittee No. 3, 1961

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