

RATES OF REACTION

REACTION RATE: The change in concentration of reactants or products per unit time.

- Can tell how quickly a reaction takes place by looking at
- How quickly reactants decrease.
 - How quickly products form.

$$\text{RATE} = \frac{\Delta C}{\Delta t}$$

mol dm⁻³ s⁻¹ s/min / hrs

Rate of reaction can also be monitored with measuring

- Δn - ΔV
- Δm - Δp

Instantaneous rate $\Rightarrow$ Tangent

COLLISION THEORY

Model that explains reaction rate as the result of particles colliding with a certain minimum energy to form products

The rate of a reaction is determined by the number of effective collisions per unit time

- Orientation of collisions
- Energy of molecules
 - $\hookrightarrow$ Temp influences E_k of molecules
- Activation energy
 - $E_k > E_a$
 - (Catalyst lower E_a)

Effective collision $\rightarrow$ results in a reaction that produces 1+ products. Reaction can only occur if collision is effective

FACTORS THAT AFFECT RATE

TEMPERATURE

$\hookrightarrow$ Directly proportional to the rate, E_k
Note: all particles have some E_k but some have more

If $T \uparrow$

- No. of molecules $\uparrow$
- Speed of molecules $\uparrow$
- More molecules have sufficient energy
- More effective collisions per unit time $\rightarrow \uparrow$ Reaction Rate

SOLID CONTACT AREA (SA)

Larger contact area $\rightarrow$ More effective collisions per unit time

CONCENTRATION OF SOLUTION

$\uparrow$ mole quantity ($\uparrow$ concentration) $\rightarrow$
 $\uparrow$ chances of an effective collision

$$C = \frac{n}{V}$$

$\uparrow n \rightarrow \uparrow C \rightarrow \uparrow V \rightarrow \uparrow n$

GAS PRESSURE

The higher the pressure of the gaseous reactants:

- The smaller the volume $\hookrightarrow$ the higher the concentration

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**Karin H. Kelder,Jagathesan
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