

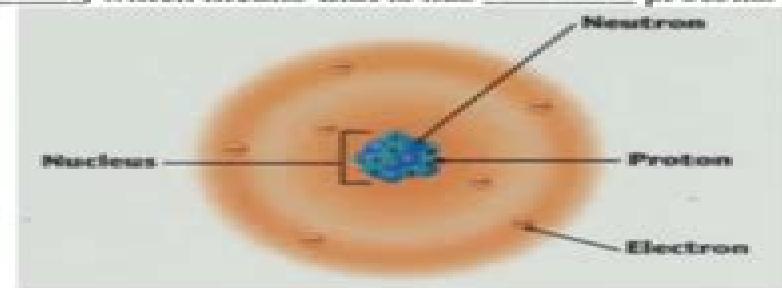
Physical Science Standards:

SPS1: Students will investigate our current understanding of the atom.

a. Examine the structure of the atom in terms of

- proton, electron, and neutron locations.
 - Proton has a charge of _____ and is located in the _____ of the atom.
 - Protons are also the elements _____ number of the identity of the element, so protons are found by looking at the atomic number
 - Neutron has a charge of _____ and is located in the _____ of the atom.
 - Neutrons are found by subtracting the _____ Number from the _____ Mass or the BIG # - SMALL #.
 - Electron has a charge of _____ and is located in the _____ of the atom.
 - Electrons are equal to the number of _____ in a neutral atom.
 - The nucleus is made up of these 2 particles _____ and _____.
 - The electrons are located in this region of the atom, known as the _____.
 - The electron cloud is divided in to energy levels or energy shells, the first level will hold _____ electrons and the second level will hold _____ electrons.
 - The valence shell or valence level is the _____ energy shell of the atom.
- atomic mass and atomic number.
 - Atomic number is equal to the number of _____ in the atom and is also the atoms identity—it's the small number of the periodic table.
 - Atomic mass is equal to the average mass of all the isotopes of a particular element—it the big number on the periodic table.
- atoms with different numbers of neutrons (isotopes).
 - Define Isotope _____.
 - Ex. Carbon-12 and Carbon-14
 - Carbon 12 has _____ P _____ N _____ E
 - Carbon 14 has _____ P _____ N _____ E
- explain the relationship of the proton number to the element's identity.
 - The _____ number is equal to the number of _____ in the atom which is how the element is identified.
 - Ex. Nitrogen has an atom number of _____, which means that it has _____ protons.

molybdenum	element name
42	atomic number number of protons (Z)
Mo	atomic symbol
95.94	atomic mass <small>A (this is an average mass)</small>



- Number of Protons = Atomic Number
- Number of Electrons = Number of Protons = Atomic Number
- Number of Neutrons = Mass Number - Atomic Number

Physical Science Standards Eoct Study Guide

Willard J. Poppy

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Table of Contents Physical Science Standards Eoct Study Guide

1. Understanding the eBook Physical Science Standards Eoct Study Guide
 - The Rise of Digital Reading Physical Science Standards Eoct Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Standards Eoct Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Standards Eoct Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Standards Eoct Study Guide
 - Personalized Recommendations
 - Physical Science Standards Eoct Study Guide User Reviews and Ratings
 - Physical Science Standards Eoct Study Guide and Bestseller Lists

5. Accessing Physical Science Standards Eoct Study Guide Free and Paid eBooks
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7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Standards Eoct Study Guide
 - Highlighting and Note-Taking Physical Science Standards Eoct Study Guide
 - Interactive Elements Physical Science Standards Eoct Study Guide
8. Staying Engaged with Physical Science Standards Eoct Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Standards Eoct Study Guide
9. Balancing eBooks and Physical Books Physical Science Standards Eoct Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Standards Eoct Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Standards Eoct Study Guide
 - Setting Reading Goals Physical Science Standards Eoct Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Standards Eoct Study Guide
 - Fact-Checking eBook Content of Physical Science Standards Eoct Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
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