

Physics

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(Chapter 11) (Thermal Properties of Matter) (Exercises)
(Class 11)

$$\text{Useful Energy, } \Delta Q = \frac{50}{100} \times 1.5 \times 10^6 = 7.5 \times 10^5 \text{ J}$$

$$\text{But } \Delta Q = mc\Delta T$$

$$\therefore \Delta T = \frac{\Delta Q}{mc} = \frac{7.5 \times 10^5}{8 \times 10^3 \times 0.91} = 103^\circ\text{C}$$

Therefore, in 2.5 minutes of drilling, the rise in the temperature of the block is 103°C .

Question 11.13:

A copper block of mass 2.5 kg is heated in a furnace to a temperature of 500°C and then placed on a large ice block. What is the maximum amount of ice that can melt? (Specific heat of copper = $0.39 \text{ J g}^{-1} \text{ K}^{-1}$; heat of fusion of water = 335 J g^{-1})

Answer 11.13:

Mass of the copper block, $m = 2.5 \text{ kg} = 2500 \text{ g}$

Rise in the temperature of the copper block, $\Delta\theta = 500^\circ\text{C}$

Specific heat of copper, $C = 0.39 \text{ J g}^{-1} \text{ }^\circ\text{C}^{-1}$

Heat of fusion of water, $L = 335 \text{ J g}^{-1}$

The maximum heat the copper block can lose, $Q = mC\Delta\theta = 2500 \times 0.39 \times 500 = 487500 \text{ J}$

Let $m_1 \text{ g}$ be the amount of ice that melts when the copper block is placed on the ice block.

The heat gained by the melted ice, $Q = m_1 L$

$$\therefore m_1 = \frac{Q}{L} = \frac{487500}{335} = 1455.22 \text{ g}$$

Hence, the maximum amount of ice that can melt is 1.45 kg.

Question 11.14:

In an experiment on the specific heat of a metal, a 0.20 kg block of the metal at 150°C is dropped in a copper calorimeter (of water equivalent 0.025 kg) containing 150 cm^3 of water at 27°C . The final temperature is 40°C . Compute the specific heat of the metal. If heat losses to the surroundings are not negligible, is your answer greater or smaller than the actual value for specific heat of the metal?

Answer 11.14:

Mass of the metal, $m = 0.20 \text{ kg} = 200 \text{ g}$

Initial temperature of the metal, $T_1 = 150^\circ\text{C}$ and Final temperature of the metal, $T_2 = 40^\circ\text{C}$

Calorimeter has water equivalent of mass, $m' = 0.025 \text{ kg} = 25 \text{ g}$

Volume of water, $V = 150 \text{ cm}^3$

Mass (M) of water at temperature $T = 27^\circ\text{C}$: $150 \times 1 = 150 \text{ g}$

Fall in the temperature of the metal: $\Delta T = T_1 - T_2 = 150 - 40 = 110^\circ\text{C}$

Specific heat of water, $C_w = 4.186 \text{ J/g/}^\circ\text{K}$ and Specific heat of the metal = C

Heat lost by the metal, $\theta = mC\Delta T \dots (i)$

Rise in the temperature of the water and calorimeter system:

$$\Delta T' = 40 - 27 = 13^\circ\text{C}$$

Heat gained by the water and calorimeter system: $\Delta\theta' = m_1 C_w \Delta T' = (M + m') C_w \Delta T' \dots (ii)$

Heat lost by the metal = Heat gained by the water and calorimeter system

$$mC\Delta T = (M + m') C_w \Delta T' \Rightarrow 200 \times C \times 110 = (150 + 25) \times 4.186 \times 13$$

$$\therefore C = \frac{175 \times 4.186 \times 13}{110 \times 200} = 0.43 \text{ J g}^{-1} \text{ K}^{-1}$$

If some heat is lost to the surroundings, then the value of C will be smaller than the actual value.

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**John Barry Cook,R. T. M. Greene,Bill
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1. Understanding the eBook Physics 1st Paper For Class11 Nctb
 - The Rise of Digital Reading Physics 1st Paper For Class11 Nctb
 - Advantages of eBooks Over Traditional Books
2. Identifying Physics 1st Paper For Class11 Nctb
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physics 1st Paper For Class11 Nctb
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physics 1st Paper For Class11 Nctb
 - Personalized Recommendations
 - Physics 1st Paper For Class11 Nctb User Reviews and Ratings
 - Physics 1st Paper For Class11 Nctb and Bestseller Lists
5. Accessing Physics 1st Paper For Class11 Nctb Free and Paid eBooks
 - Physics 1st Paper For Class11 Nctb Public Domain eBooks
 - Physics 1st Paper For Class11 Nctb eBook Subscription Services

- Physics 1st Paper For Class11 Nctb Budget-Friendly Options
- 6. Navigating Physics 1st Paper For Class11 Nctb eBook Formats
 - ePub, PDF, MOBI, and More
 - Physics 1st Paper For Class11 Nctb Compatibility with Devices
 - Physics 1st Paper For Class11 Nctb Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physics 1st Paper For Class11 Nctb
 - Highlighting and Note-Taking Physics 1st Paper For Class11 Nctb
 - Interactive Elements Physics 1st Paper For Class11 Nctb
- 8. Staying Engaged with Physics 1st Paper For Class11 Nctb
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physics 1st Paper For Class11 Nctb
- 9. Balancing eBooks and Physical Books Physics 1st Paper For Class11 Nctb
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physics 1st Paper For Class11 Nctb
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physics 1st Paper For Class11 Nctb
 - Setting Reading Goals Physics 1st Paper For Class11 Nctb
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physics 1st Paper For Class11 Nctb
 - Fact-Checking eBook Content of Physics 1st Paper For Class11 Nctb
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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