

**Activity 11 :Chemical Formula for ionic compounds:**

Complete the table below :

Cation	Formula	Anion	Formula
Hydrogen ion	$H^+$	Fluoride ion	$F^-$
Lithium ion		Chloride ion	
Sodium ion		Bromide ion	
Potassium ion		Iodide ion	
Magnesium ion		Hydroxide ion	
Calcium ion	$Ca^{2+}$	Nitrate ion	
Barium ion	$Ba^{2+}$	Manganate(VII) ion	
Copper(II) ion		Ethanoate ion	$CH_3COO^-$
Iron(II) ion			$O^{2-}$
Iron (III) ion		Sulphate ion	
Lead (II) ion		Sulphide ion	$S^{2-}$
Zinc ion		Carbonate ion	
Chromium (III) ion		Dichromate (VI) ion	$Cr_2O_7^{2-}$
Aluminium ion	$Al^{3+}$		$PO_4^{3-}$
Ammonium ion		Chromate (VI) ion	

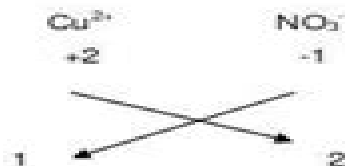
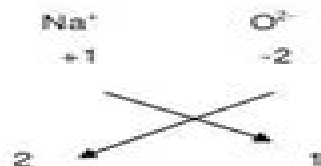
## Activity 12

a) Chemical formula of an ionic compound comprising of the ions  $X^{m+}$  and  $Y^{n-}$  is constructed

by exchanging the charges of each element. The formula obtained will  $X_nY_m$ .

Example : Sodium oxide

Copper (II) nitrate



# Writing Ionic Chemical Formulas Answers

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