



# CHEMISTRY

## CHEAT SHEET

### WRITING IONIC FORMULAS

#### IONIC COMPOUNDS CHEAT SHEET

##### WRITING IONIC FORMULAS

**How to Write Ionic Formulas**

1. Write the symbols for the cation and anion, including their charges.

2. Write the cation symbol first, followed by the anion symbol.

3. Use the criss-cross method to determine the subscripts for each ion.

4. Write the formula with the subscripts.

5. Simplify the formula if possible.

Ion	Charge	Symbol
Hydrogen	+1	H <sup>+</sup>
Sodium	+1	Na <sup>+</sup>
Potassium	+1	K <sup>+</sup>
Ammonium	+1	NH <sub>4</sub> <sup>+</sup>
Lithium	+1	Li <sup>+</sup>
Calcium	+2	Ca <sup>2+</sup>
Strontium	+2	Sr <sup>2+</sup>
Barium	+2	Ba <sup>2+</sup>
Magnesium	+2	Mg <sup>2+</sup>
Zinc	+2	Zn <sup>2+</sup>
Cadmium	+2	Cd <sup>2+</sup>
Mercury (I)	+1	Hg <sub>2</sub> <sup>+</sup>
Mercury (II)	+2	Hg <sup>2+</sup>
Copper (I)	+1	Cu <sup>+</sup>
Copper (II)	+2	Cu <sup>2+</sup>
Nickel	+2	Ni <sup>2+</sup>
Cobalt	+2	Co <sup>2+</sup>
Iron	+2	Fe <sup>2+</sup>
Iron	+3	Fe <sup>3+</sup>
Chromium	+2	Cr <sup>2+</sup>
Chromium	+3	Cr <sup>3+</sup>
Manganese	+2	Mn <sup>2+</sup>
Aluminum	+3	Al <sup>3+</sup>
Gallium	+3	Ga <sup>3+</sup>
Indium	+3	In <sup>3+</sup>
Thallium	+3	Tl <sup>3+</sup>
Lead	+2	Pb <sup>2+</sup>
Lead	+4	Pb <sup>4+</sup>
Bismuth	+3	Bi <sup>3+</sup>
Antimony	+3	Sb <sup>3+</sup>
Arsenic	+3	As <sup>3+</sup>
Vanadium	+2	V <sup>2+</sup>
Vanadium	+3	V <sup>3+</sup>
Vanadium	+4	V <sup>4+</sup>
Vanadium	+5	V <sup>5+</sup>
Chromium	+6	Cr <sup>6+</sup>
Manganese	+7	Mn <sup>7+</sup>
Fluoride	-1	F <sup>-</sup>
Chloride	-1	Cl <sup>-</sup>
Bromide	-1	Br <sup>-</sup>
Iodide	-1	I <sup>-</sup>
Nitride	-3	N <sup>3-</sup>
Oxide	-2	O <sup>2-</sup>
Sulfide	-2	S <sup>2-</sup>
Selenide	-2	Se <sup>2-</sup>
Telluride	-2	Te <sup>2-</sup>
Carbonate	-2	CO <sub>3</sub> <sup>2-</sup>
Bicarbonate	-1	HCO <sub>3</sub> <sup>-</sup>
Hydroxide	-1	OH <sup>-</sup>
Perchlorate	-1	ClO <sub>4</sub> <sup>-</sup>
Chlorate	-1	ClO <sub>3</sub> <sup>-</sup>
Sulfate	-2	SO <sub>4</sub> <sup>2-</sup>
Bisulfate	-1	HSO <sub>4</sub> <sup>-</sup>
Phosphate	-3	PO <sub>4</sub> <sup>3-</sup>
Biphosphate	-2	H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>
Dihydrogen phosphate	-1	H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>
Hydrogen phosphate	-2	HPO <sub>4</sub> <sup>2-</sup>
Acetate	-1	CH <sub>3</sub> COO <sup>-</sup>
Formate	-1	HCOO <sup>-</sup>
Oxalate	-2	C <sub>2</sub> O <sub>4</sub> <sup>2-</sup>
Malonate	-2	C <sub>3</sub> O <sub>4</sub> <sup>2-</sup>
Succinate	-2	C <sub>4</sub> O <sub>4</sub> <sup>2-</sup>
Fumarate	-2	C <sub>4</sub> O <sub>4</sub> <sup>2-</sup>
Maleate	-2	C <sub>4</sub> O <sub>4</sub> <sup>2-</sup>
Citrate	-3	C <sub>6</sub> H <sub>5</sub> O <sub>7</sub> <sup>3-</sup>
Ascorbate	-2	C <sub>6</sub> H <sub>7</sub> O <sub>6</sub> <sup>2-</sup>
Lactate	-1	C <sub>3</sub> H <sub>5</sub> O <sub>2</sub> <sup>-</sup>
Pyruvate	-1	C <sub>3</sub> H <sub>3</sub> O <sub>3</sub> <sup>-</sup>
Glucuronate	-2	C <sub>6</sub> H <sub>7</sub> O <sub>6</sub> <sup>2-</sup>
Galacturonate	-2	C <sub>6</sub> H <sub>7</sub> O <sub>6</sub> <sup>2-</sup>
Mannuronate	-2	C <sub>6</sub> H <sub>7</sub> O <sub>6</sub> <sup>2-</sup>
Xylofuranose	-2	C <sub>5</sub> H <sub>7</sub> O <sub>5</sub> <sup>2-</sup>
Arabinofuranose	-2	C <sub>5</sub> H <sub>7</sub> O <sub>5</sub> <sup>2-</sup>
Ribofuranose	-2	C <sub>5</sub> H <sub>7</sub> O <sub>5</sub> <sup>2-</sup>
Deoxyribose	-2	C <sub>5</sub> H <sub>10</sub> O <sub>4</sub> <sup>2-</sup>
Glucose	-2	C <sub>6</sub> H <sub>12</sub> O <sub>6</sub> <sup>2-</sup>
Fructose	-2	C <sub>6</sub> H <sub>12</sub> O <sub>6</sub> <sup>2-</sup>
Saccharose	-2	C <sub>12</sub> H <sub>22</sub> O <sub>11</sub> <sup>2-</sup>
Lactose	-2	C <sub>12</sub> H <sub>22</sub> O <sub>11</sub> <sup>2-</sup>
Maltose	-2	C <sub>12</sub> H <sub>22</sub> O <sub>11</sub> <sup>2-</sup>
Starch	-2	(C <sub>6</sub> H <sub>10</sub> O <sub>5</sub> ) <sub>n</sub> <sup>2-</sup>
Cellulose	-2	(C <sub>6</sub> H <sub>10</sub> O <sub>5</sub> ) <sub>n</sub> <sup>2-</sup>
Chitin	-2	(C <sub>5</sub> H <sub>7</sub> NO <sub>2</sub> ) <sub>n</sub> <sup>2-</sup>
Chitosan	-2	(C <sub>5</sub> H <sub>9</sub> NO <sub>2</sub> ) <sub>n</sub> <sup>2-</sup>
Alginate	-2	(C <sub>5</sub> H <sub>7</sub> NO <sub>2</sub> ) <sub>n</sub> <sup>2-</sup>
Agar	-2	(C <sub>5</sub> H <sub>7</sub> NO <sub>2</sub> ) <sub>n</sub> <sup>2-</sup>
Carrageenan	-2	(C <sub>5</sub> H <sub>7</sub> NO <sub>2</sub> ) <sub>n</sub> <sup>2-</sup>
Pectin	-2	(C <sub>5</sub> H <sub>7</sub> NO <sub>2</sub> ) <sub>n</sub> <sup>2-</sup>
Dextran	-2	(C <sub>6</sub> H <sub>10</sub> O <sub>5</sub> ) <sub>n</sub> <sup>2-</sup>
Inulin	-2	(C <sub>6</sub> H <sub>10</sub> O <sub>5</sub> ) <sub>n</sub> <sup>2-</sup>
Starch	-2	(C <sub>6</sub> H <sub>10</sub> O <sub>5</sub> ) <sub>n</sub> <sup>2-</sup>
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Cellulose	-2	(C <sub>6</sub> H <sub>10</sub> O <sub>5</sub> ) <sub>n</sub> <sup>2-</sup>
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Cellulose	-2	(C <sub>6</sub> H <sub>10</sub> O <sub>5</sub> ) <sub>n</sub> <sup>2-</sup>
Chitin	-2	(C <sub>5</sub> H <sub>7</sub> NO <sub>2</sub> ) <sub>n</sub> <sup>2-</sup>
Chitosan	-2	(C <sub>5</sub> H <sub>9</sub> NO <sub>2</sub> ) <sub>n</sub> <sup>2-</sup>
Alginate	-2	(C <sub>5</sub> H <sub>7</sub> NO <sub>2</sub> ) <sub>n</sub> <sup>2-</sup>
Agar	-2	(C <sub>5</sub> H <sub>7</sub> NO <sub>2</sub> ) <sub>n</sub> <sup>2-</sup>
Carrageenan	-2	(C <sub>5</sub> H <sub>7</sub> NO <sub>2</sub> ) <sub>n</sub> <sup>2-</sup>
Pectin	-2	(C <sub>5</sub> H <sub>7</sub> NO <sub>2</sub> ) <sub>n</sub> <sup>2-</sup>
Dextran	-2	(C <sub>6</sub> H <sub>10</sub> O <sub>5</sub> ) <sub>n</sub> <sup>2-</sup>
Inulin	-2	(C <sub>6</sub> H <sub>10</sub> O <sub>5</sub> ) <sub>n</sub> <sup>2-</sup>
Starch	-2	(C <sub>6</sub> H <sub>10</sub> O <sub>5</sub> ) <sub>n</sub> <sup>2-</sup>
Cellulose	-2	(C <sub>6</sub> H <sub>10</sub> O <sub>5</sub> ) <sub>n</sub> <sup>2-</sup>
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# Writing Ionic Formulas Study Guide

**Martha Mackin**



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**Chemistry, Student Study Guide** James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement *Study Guide to Accompany Basics for Chemistry* Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students **Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone s Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each

chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

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Study Guide to Accompany Calculus for the Management, Life, and Social Sciences Clyde Metz, 1984-01-01 Study Guide to Accompany Calculus for the Management Life and Social Sciences *Study Guide with Selected Solutions* Karen Timberlake, 2002-08-06 Written by the author the Study Guide is keyed to the learning goals in the text and designed to promote active learning through a variety of exercises with answers and mastery exams Also contains complete solutions to odd numbered problems

**General Chemistry Study Guide Sixth Edition** Darrell D. Ebbing, 1999 **Study Guide** Darrell D. Ebbing, Susan M. Schelble, 1998 **Barron's Chemistry Practice Plus: 400+ Online Questions and Quick Study Review** Barron's Educational Series, Mark Kernion, Joseph A. Mascetta, 2022-07-05 Barron's Chemistry Practice Plus features more than 400 online practice questions and a concise review guide that covers the basics of Chemistry Inside you'll find concise review on the basics of Chemistry an excellent resource for students who want a quick review of the most important topics access to 400 online questions arranged by topic for customized practice online practice includes answer explanations with expert advice for all questions plus scoring to track your progress This essential guide is the perfect practice supplement for students and teachers

Fundamentals of Chemistry, Study Guide James E. Brady, John R. Holum, 1988-04-20 This Third Edition of the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up to date treatments of important and current topics Emphasizes descriptive chemistry chemical reactions and properties while maintaining a solid treatment of chemical principles Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises Incorporates many pedagogical aids each chapter begins with a brief table of contents and each section begins with a preview of topics covered Chapters include frequent margin comments figures and photographs

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