

### Balancing Chemical Equations Worksheet

1. 2  $\text{H}_2$  + 1  $\text{O}_2 \rightarrow$  2  $\text{H}_2\text{O}$
2. 1  $\text{N}_2$  + 3  $\text{H}_2 \rightarrow$  2  $\text{NH}_3$
3. 1  $\text{S}_8$  + 12  $\text{O}_2 \rightarrow$  8  $\text{SO}_3$
4. 2  $\text{N}_2$  + 1  $\text{O}_2 \rightarrow$  2  $\text{N}_2\text{O}$
5. 2  $\text{HgO} \rightarrow$  2  $\text{Hg}$  + 1  $\text{O}_2$
6. 6  $\text{CO}_2$  + 6  $\text{H}_2\text{O} \rightarrow$  1  $\text{C}_6\text{H}_{12}\text{O}_6$  + 6  $\text{O}_2$
7. 1  $\text{Zn}$  + 2  $\text{HCl} \rightarrow$  1  $\text{ZnCl}_2$  + 1  $\text{H}_2$
8. 1  $\text{SiCl}_4$  + 4  $\text{H}_2\text{O} \rightarrow$  1  $\text{H}_4\text{SiO}_4$  + 4  $\text{HCl}$
9. 2  $\text{Na}$  + 2  $\text{H}_2\text{O} \rightarrow$  2  $\text{NaOH}$  + 1  $\text{H}_2$
10. 2  $\text{H}_3\text{PO}_4 \rightarrow$  1  $\text{H}_4\text{P}_2\text{O}_7$  + 1  $\text{H}_2\text{O}$
11. 1  $\text{C}_{10}\text{H}_{16}$  + 8  $\text{Cl}_2 \rightarrow$  10  $\text{C}$  + 16  $\text{HCl}$
12. 1  $\text{CO}_2$  + 2  $\text{NH}_3 \rightarrow$  1  $\text{OC}(\text{NH}_2)_2$  + 1  $\text{H}_2\text{O}$
13. 4  $\text{Si}_2\text{H}_3$  + 17  $\text{O}_2 \rightarrow$  8  $\text{SiO}_2$  + 6  $\text{H}_2\text{O}_3$
14. 2  $\text{Al}(\text{OH})_3$  + 3  $\text{H}_2\text{SO}_4 \rightarrow$  1  $\text{Al}_2(\text{SO}_4)_3$  + 6  $\text{H}_2\text{O}$
15. 4  $\text{Fe}$  + 3  $\text{O}_2 \rightarrow$  2  $\text{Fe}_2\text{O}_3$
16. 1  $\text{Fe}_2(\text{SO}_4)_3$  + 6  $\text{KOH} \rightarrow$  3  $\text{K}_2\text{SO}_4$  + 2  $\text{Fe}(\text{OH})_3$
17. 2  $\text{C}_7\text{H}_6\text{O}_2$  + 15  $\text{O}_2 \rightarrow$  14  $\text{CO}_2$  + 6  $\text{H}_2\text{O}$
18. 1  $\text{H}_2\text{SO}_4$  + 8  $\text{HI} \rightarrow$  1  $\text{H}_2\text{S}$  + 4  $\text{I}_2$  + 4  $\text{H}_2\text{O}$
19. 4  $\text{FeS}_2$  + 11  $\text{O}_2 \rightarrow$  2  $\text{Fe}_2\text{O}_3$  + 8  $\text{SO}_2$
20. 2  $\text{Al}$  + 3  $\text{FeO} \rightarrow$  1  $\text{Al}_2\text{O}_3$  + 3  $\text{Fe}$
21. 1  $\text{Fe}_2\text{O}_3$  + 3  $\text{H}_2 \rightarrow$  2  $\text{Fe}$  + 3  $\text{H}_2\text{O}$
22. 1  $\text{Na}_2\text{CO}_3$  + 2  $\text{HCl} \rightarrow$  2  $\text{NaCl}$  + 1  $\text{H}_2\text{O}$  + 1  $\text{CO}_2$
23. 2  $\text{K}$  + 1  $\text{Br}_2 \rightarrow$  2  $\text{KBr}$
24. 1  $\text{C}_7\text{H}_{16}$  + 11  $\text{O}_2 \rightarrow$  7  $\text{CO}_2$  + 8  $\text{H}_2\text{O}$
25. 1  $\text{P}_4$  + 5  $\text{O}_2 \rightarrow$  2  $\text{P}_2\text{O}_5$

# Ws 4 1 Balancing Equations Answers

**Victor S Krimsley, Darold E  
Skerritt, Beverly B Harrison**



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class aid studying and improve reading writing and comprehension skills It can also generate discussion topics facilitate collaborative story creation and assist with writing assignments By the end of this guide you will have a deep understanding of how to effectively integrate ChatGPT into your teaching practice enabling you to create dynamic and engaging learning experiences for your students You can copy the prompts into ChatGPT to generate the products that you want We hope that this e book will inspire you to embrace the power of AI and become a part of the exciting future of education technology Let s embark on this learning journey together and unlock the potential of AI in education

**I/M&w/Tsts Intro Chem** Victor S Krimsley,Darold E Skeritt,Beverly B Harrison,1986 **Journal of Engineering for Gas Turbines and Power** ,2009

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