

**Balancing a Redox Equation for an Acidic Solution  
Using the Oxidation Number Method**  
(Student textbook page 615)

35. Use the oxidation number method to balance the following equation for the combustion of carbon disulfide:



**What Is Required?**

You must balance an equation for the combustion of carbon disulfide, using the oxidation number method.

**What Is Given?**

You are given the reactants, carbon disulfide and oxygen, and the products, carbon dioxide and sulfur dioxide.

You are given rules for assigning oxidation numbers in **Table 9.3** on page 604 of the student textbook.

Plan Your Strategy	Act on Your Strategy
<b>Step 1</b> Assign an oxidation number to each atom in the equation and determine whether it is a redox reaction.	$\overset{+4 \ -2}{\text{CS}_2}(\text{g}) + \overset{0}{\text{O}_2}(\text{g}) \rightarrow \overset{+4 \ -2}{\text{CO}_2}(\text{g}) + \overset{+4 \ -2}{\text{SO}_2}(\text{g})$ <p>The oxidation numbers of sulfur and oxygen change during the reaction. It is a redox reaction.</p>
<b>Step 2</b> Identify the atom or atoms that undergo an increase in oxidation number and the atom or atoms that undergo a decrease in oxidation number.	<p>Each of two sulfur atoms lost 6 electrons and each of two oxygen atoms gained 2 electrons. So the <math>\text{O}_2</math> lost 12 electrons and the <math>\text{CS}_2</math> gained 4 electrons.</p> $  \begin{array}{c}  \text{1}(-12\text{e}^-) = -12\text{e}^- \\  \downarrow \\  \overset{+4 \ -2}{\text{CS}_2}(\text{g}) + \overset{0}{\text{O}_2}(\text{g}) \rightarrow \overset{+4 \ -2}{\text{CO}_2}(\text{g}) + \overset{+4 \ -2}{\text{SO}_2}(\text{g}) \\  \uparrow \qquad \qquad \qquad \uparrow \\  \text{3}(+4\text{e}^-) = +12\text{e}^-  \end{array}  $

# Solutions Manual Redox Reactions

**James E. Brady, Joel W. Russell, John R.  
Holum**



## **Solutions Manual Redox Reactions:**

**Solutions Manual to Accompany Inorganic Chemistry** Alen Hadzovic, 2018 As you master each chapter in Inorganic Chemistry having detailed solutions handy allows you to confirm your answers and develop your ability to think through the problem solving process

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many examples from biology and nanoscience In this way the book helps the student to see the frontier of modern physical chemistry developments The book begins with a discussion of wave particle duality and proceeds systematically to more complex chemical systems in order to relate the story of physical chemistry in an intellectually coherent manner The topics are organized to correspond with those typically given in each of a two course semester sequence The first 13 chapters present quantum mechanics and spectroscopy to describe and predict the structure of matter atoms molecules and solids Chapters 14 to 29 present statistical thermodynamics and kinetics and applies their principles to understanding equilibria chemical transformations macromolecular properties and supramolecular machines Each chapter of the book begins with a simplified view of a topic and evolves to more rigorous description in order to provide the student and instructor flexibility to choose the level of rigor and detail that suits them best The textbook treats important new directions in physical chemistry research including chapters on macromolecules principles of interfaces and films for organizing matter and supramolecular machines as well as including discussions of modern nanoscience spectroscopy and reaction dynamics throughout the text

Student Solutions Manual to Accompany Atkins' Physical Chemistry Peter Bolgar, Haydn Lloyd, James Keeler, Aimee North, Vladimiras Oleinikovas, Stephanie Smith, 2018 The Student Solutions Manual to accompany Atkins Physical Chemistry 11th Edition provides full worked solutions to the a exercises and the odd numbered discussion questions and problems presented in the parent book The manual is intended for students and provides helpful comments and friendly advice to aid understanding

**Student Solutions Manual for Chang's Chemistry** Philip C. Keller, Raymond Chang, Jill L. Keller, 1988

**Fundamentals of Environmental Sampling and Analysis** Chunlong Zhang, 2024-07-03 This is the Solutions Manual to accompany Fundamentals of Environmental Sampling and Analysis Second Edition It provides solutions to the exercises and problems found in the main volume This book introduces a comprehensive overview on the fundamentals and applications of environmental sampling and analysis for students in environmental science and engineering as well as environmental professionals involved in sampling and analytical work The book details fundamentals of sampling selection of standard methods QA QC sample preparation chemical and instrumental principles and method applications to various contaminants in environmental matrices air water soil waste and biological samples The book gives an integrated introduction to sampling and analysis both are essential to quality environmental data For example contrary to other books that introduce a specific area of sampling and analysis this text provides a balanced mix of field sampling and laboratory analysis essential knowledge in chemistry statistics hydrology regulations wet chemical methods for conventional chemicals as well as various modern instrumental techniques for contaminants of emerging concerns The new edition adds three standalone chapters regarding the basics of analytical and organic chemistry environmental data analysis mass spectrometry and other significant amounts of new materials such as time integrated passive sampling incremental sampling green sample preparation Raman spectroscopy chiral separation and non target analysis In addition the second edition provides more examples visual aids

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**Study Guide and Partial Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry** Susan McMurry, 1996 This internationally acclaimed detective series is just the thing for lovers of those Number One Ladies looking for a darker more realistic view of Botswana Sue Baker Publishing News

**Chemistry in Context - Laboratory Manual** Graham Hill, John Stranger Holman, John Holman, 2001 The laboratory manual and study guide supports your teaching with a broad range of practicals emphasising safety and risk assessment It is an essential companion to Chemistry in Context and can also be used alongside other Advanced Chemistry books It offers practicals with detailed instructions for open ended investigations and opportunities for assessed practical work in the four skill areas of planning implementing analysing and evaluating

**Chemistry, Student Solutions Manual** James N. Spencer, George M. Bodner, Lyman H. Rickard, 2008-01-09 This innovative text provides a 15 chapter introduction to the fundamental concepts of chemistry The material is then supplemented by special topics at the end of each chapter

**A Manual for the Chemical Analysis of Metals ,**

**Cehmistry Textbook for College and University USA** Ibrahim Sikder, 2023-06-04 Cehmistry Textbook USA

Chemistry, Student Solutions Manual James E. Brady, Joel W. Russell, John R. Holum, 2000-01-26 The images on the cover call attention to the relationship between macro observations and the intimate structure of chemical substances and the changes both chemical and physical that they undergo

**Fireworks** One of the ingredients is phosphorus a molecular form of which is believed to consist of linked tetrahedra of phosphorus atoms The chemical reaction of phosphorus with oxygen is partly responsible for the spectacular show of light

**Carbon** The element is found in several forms including the familiar diamond and another recently discovered sooty substance that consists of soccer ball shaped molecules often referred to as buckeyballs Diamond is not the most stable form of carbon and is created from other forms of carbon at high temperatures and pressures deep within the earth

**Acetylene torch** Cutting steel is possible because of the intense heat generated by the chemical reaction of acetylene with oxygen a reaction between molecules of  $C_2H_2$  and  $O_2$  to give  $CO_2$  and  $H_2O$

**Hot air balloon** The air that helps it rise is heated by the combustion of molecules of propane each composed of three carbon and eight hydrogen atoms

**Stormy weather** The evaporation of water serves to store energy provided by the sun Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet

**Handbook of Air Pollution Analysis** Roy M. Harrison, 2012-12-06 The contributing authors have been chosen because of their depth of knowledge and experience in air pollution work and we are confident that this is reflected in a Handbook which will find very wide application wherever air pollution analysis is practised Roy M Harrison Roger Perry February 1985 Readers are recommended to follow all the usual laboratory safety precautions While care has been taken to ensure that the information in this book is correct neither the authors nor the publisher can accept responsibility for any outcome of the application of methods and procedures outlined in this book

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